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Gas-solid carbonation as a possible source of carbonates in cold planetary environments

Alexandre Garenne, German Montes-Hernandez, Pierre Beck, Bernard
Schmitt, Olivier Brissaud, Antoine Pommerol

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26 **Abstract**

27 Carbonates are abundant sedimentary minerals at the surface and sub-surface of the Earth and
28 they have been proposed as tracers of liquid water in extraterrestrial environments. Their
29 formation mechanism is since generally associated with aqueous alteration processes. Recently,
30 carbonate minerals have been discovered on Mars' surface by different orbital or rover missions.
31 In particular, the phoenix mission has measured from 1 to 5% of calcium carbonate (calcite type)
32 within the soil (Smith P.H. et al., 2009). These occurrences have been reported in area where the
33 relative humidity is significantly high (Boynton et al., 2009). The small concentration of
34 carbonates suggests an alternative process on mineral grain surfaces (as suggested by Shaheen et
35 al., 2010) than carbonation in aqueous conditions. Such an observation could rather point toward
36 a possible formation mechanism by dust-gas reaction under current Martian conditions. To
37 understand the mechanism of carbonate formation under conditions relevant to current Martian
38 atmosphere and surface, we designed an experimental setup consisting of an infrared microscope
39 coupled to a cryogenic reaction cell (IR-CryoCell setup). Three different mineral precursors of
40 carbonates (Ca and Mg hydroxides, and a hydrated Ca silicate formed from Ca_2SiO_4), low
41 temperature (from -10 to $+30^\circ\text{C}$), and reduced CO_2 pressure (from 100 to 2000 mbar) were
42 utilized to investigate the mechanism of gas-solid carbonation at mineral surfaces. These mineral
43 materials are crucial precursors to form Ca and Mg carbonates in humid environments ($0 <$
44 $\text{relative humidity} < 100\%$) at dust- CO_2 or dust-water ice- CO_2 interfaces. Our results reveal a
45 significant and fast carbonation process for Ca hydroxide and hydrated Ca silicate. Conversely,
46 only a moderate carbonation is observed for the Mg hydroxide. These results suggest that gas-
47 solid carbonation process or carbonate formation at the dust-water ice- CO_2 interfaces could be a

48 currently active Mars' surface process. To the best of our knowledge, we report for the first time
49 that calcium carbonate can be formed at a negative temperature (-10°C) via gas-solid carbonation
50 of Ca hydroxide. We note that the carbonation process at low temperature (<0°C) described in the
51 present study could also have important implications on the dust-water ice-CO₂ interactions in
52 cold terrestrial environments (e.g. Antarctic).

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59 **Keywords:** Carbonates; Gas-solid carbonation; Mars; Low temperature; Infrared Microscopy; Ca
60 and Mg Hydroxides.

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66 **1. Introduction**

67 The biotic and abiotic (i.e. chemical) formation of carbonates plays a crucial role in the
68 global carbon cycle on Earth. In addition, carbonate minerals often sequester various trace
69 elements (actinides and lanthanides), metalloids, and heavy metals, and thus control in part their
70 global cycling (e.g. Paquette and Reeder, 1995; Stumm and Morgan, 1995; Sigg et al., 2000). In
71 general, carbonate minerals can be formed in natural or artificial environments by three different
72 mechanisms (e.g. Montes-Hernandez et al., 2010a): (1) aqueous nucleation-growth in
73 homogeneous or heterogeneous systems (aqueous conditions), for example, the chemical or
74 biogenic formation of carbonates in lakes, oceans, CO₂ storage sites, natural caves; (2) gas-solid
75 carbonation of alkaline minerals (fine particles) in the presence of adsorbed water (water
76 humidity conditions, $0 < \text{water activity} < 1$), for example carbonate formation in water-
77 unsaturated soils, in terrestrial or extraterrestrial aerosols (Shaheen et al., 2010). This water has
78 an important role in the surface chemistry of minerals as was shown by Galhotra et al., (2009)
79 and Baltrusaitis and Grassian (2005) with zeolites and iron oxide surfaces; (3) dry gas-solid
80 carbonation of granular/porous materials (dry conditions, water activity ≈ 0), for example, the
81 industrial mineralization, recovery or capture of CO₂ at high temperatures in presence of alkaline
82 binary oxides (CaO, MgO) or metastable, nanoparticle alkaline silicates (Montes-Hernandez et
83 al., 2012).

84 In the Planetary Sciences context, carbonates are generally considered as indicators of
85 aqueous alteration processes (Bandfield et al., 2003; Milliken and Rivkin, 2009; Boynton et al.,
86 2009; Ehlmann et al., 2008; Michalski and Niles, 2010). In the case of Mars, huge deposits of
87 surface carbonates remained undetected for a long period, and their suspected absence was used

88 to constrain the chemistry of a putative Martian ocean (Fairén et al., 2004). Evidences are now
89 growing for the presence of carbonates at the surface of the red planet, which include
90 observations of carbonate-rich outcrops (Ehlmann et al., 2008; Michalski and Niles, 2010) as
91 well as carbonates within the Martian dust (Bandfield et al., 2003; Boynton et al., 2009). The
92 aqueous alteration of mafic rocks in the presence of CO₂ is certainly an efficient mechanism for
93 carbonate synthesis, an alternative pathway of carbonate synthesis exists, which does not require
94 the presence of liquid water. This pathway involves reaction of a mineral substrate with CO₂ in
95 the presence of chemisorbed water (few angstroms to few nm thick layers), and was recently
96 tested and observed for terrestrial aerosols (Shaheen et al., 2010).

97 Here, we report on an experimental study of the kinetic of carbonation in liquid-water free
98 environment. We designed novel, state of the art experimental setup (IR-CryoCell) to investigate
99 the *in-situ* gas-solid carbonation (i.e. resolved in time), for temperature and pressure conditions
100 relevant to Mars. We studied carbonate synthesis starting from Ca and Mg hydroxides and an
101 amorphous silicate (synthesized from Ca₂SiO₄), at low temperature (from -10 to +30°C) and at
102 low CO₂ pressure (from 100 to 2000 mbar). These starting materials are known precursors to
103 form respective Ca and Mg carbonates in humid environments at dust-CO₂ or dust-water ice-CO₂
104 interfaces, at least under « terrestrial » conditions. They also can be expected to occur at the
105 surface of Mars and some asteroids (Mg hydroxide has been described on Ceres). We report here
106 laboratory experiments on gas-solid carbonation process at low temperature (<0°C), which
107 provides new insights on conditions for carbonate formation. We will show that gas-solid
108 carbonation can occur below the water frost point (at terrestrial atmospheric pressure), with
109 significant implications on the dust/water-ice/CO₂ interactions in cold environments.

110 **2. Materials and methods**

111 The experiments were performed using three different materials, Ca, Mg hydroxide and a Ca
112 silicate hydrate. CO₂ is known to react with surface of CaO and MgO by adsorption (Ochs et al.,
113 1998a; Ochs et al., 1998b) and produce carbonates as well the importance of OH groups to water
114 adsorption on surfaces (Yamamoto et al., 2008). These substrates were chosen to mimic natural
115 conditions and to catalyze reaction as their surfaces are terminated by OH groups: i) in order to
116 form Ca-Mg carbonate by reaction with CO₂, a Ca and Mg source is needed; ii) the presence of
117 hydroxyl groups in the starting material was requested to permit auto-catalysis of the reaction
118 (Montes-Hernandez et al., 2010a); iii) the material had to be geologically relevant.

119 Brucite has not been detected on the Martian surface. However, various types of phyllosilicates
120 have been now described over the planet, that are interpreted as aqueous alteration products of
121 mafic rocks (see the recent review by Ehlmann et al., 2011). Such aqueous alteration processes
122 can be accompanied by the production of brucite (Evans, 2008). Identification of brucite by its
123 spectral properties is difficult since no diagnostic band is present in the NIR, with the exception
124 of the 2.7 micrometer feature ubiquitous to almost all -OH bearing phases. Brucite has been
125 diagnosed on some asteroids from observations in the mid-IR (together with carbonate). It is the
126 case of the largest main-belt object, Ceres. In addition, MgO has been proposed as a condensation
127 product in some solar nebula models, which should readily transform to brucite in the presence of
128 gaseous water or humidity (Gail and Sedlmayr, 1999).

129 Portlandite has not been reported on Mars either. On Earth, it is almost always found in
130 association with calcium carbonates, and is very difficult to observe due to its high reactivity with
131 CO₂. We chose to study portlandite because of its high catalytic reactivity which enabled to

132 provide kinetic measurements under some hours. In addition, it is a structural analog to brucite
133 and a number of X-(OH)₂ type hydroxide compounds (where X=Ni, Co, Fe, Mn, Cd). CaO has
134 also been proposed as an intermediate compound by Shaheen et al. (2010) to explain the
135 formation of calcium carbonate on Mars, which could readily transform to portlandite in the
136 presence of gaseous H₂O or humidity.

137 Finally, we used an amorphous calcium silicate hydrate synthesized from larnite (Ca₂SiO₄). This
138 material was chosen to represent an amorphous volcanic material. Volcanic activity has been
139 widespread on Mars, and volcanoclastic deposits have been described (Ehlmann et al., 2011). We
140 decided to use a pure calcium amorphous silicate (rather than a basaltic glass), in order to
141 simplify the chemistry of the system. However, one might expect a more complex chemistry for
142 Martian volcanic glasses. Our approach might appear too simplistic, but might provide grounds
143 for understanding more complex chemistries.

144

145 2.1. Materials

146 **Portlandite:** Calcium hydroxide Ca(OH)₂ was provided by Sigma-Aldrich with 96% chemical
147 purity (about 3% of CaCO₃) and 1% of other impurities. This material is characterized by platy
148 nanoparticles (sheet forms) forming micrometric aggregates with high porosity and/or high
149 specific surface area (15 m²/g). Its infrared spectrum has revealed a small amount of adsorbed
150 water at atmospheric conditions, around 0,01gH₂O/gCa(OH)₂ determined by TGA. The
151 portlandite sample was used without any physicochemical treatment.

152 **Brucite:** Magnesium hydroxide Mg(OH)₂ was provided by Fisher Scientific (UK). This material

153 is characterized by platy hexagonal microparticles. A small amount of adsorbed water at
154 atmospheric conditions was detected by infrared spectroscopy. The brucite sample was crushed in
155 a mortar before use.

156 **Amorphous calcium silicate hydrate:** This material was synthesized from synthetic larnite
157 mineral (Ca_2SiO_4) by using a simple acidic treatment (2M HCl solution) at room lab temperature
158 during 15 minutes. Then, consecutive dilutions with demineralized water were carried out until
159 pH equal to 3. Finally, the solid was separated from the solution by centrifugation (10 minutes at
160 12000 rpm) and dried directly in the centrifugation flasks at 80°C for 48h. The larnite synthetic
161 mineral was provided by A. Santos and it was synthesized as reported in Santos et al. (2009).

162 **Carbon dioxide:** Carbon dioxide CO_2 was provided by Linde Gas S.A. with 99.995% of
163 chemical purity. This gas was directly injected in the cryogenic reaction cell without any
164 treatment or purification.

165 *2.2. Infrared microscope*

166 An infrared microscope (BRUKER HYPERION 3000) coupled with a cryogenic cell (designed
167 and built at IPAG) was used to obtain infrared spectra in transmission mode. The IR beam was
168 focused through a 15x objective and the typical size of the spot on the sample was around 50×50
169 μm^2 . The spectral resolution was 4cm^{-1} and the spectra were recorded in transmission mode
170 between 4000cm^{-1} and 700cm^{-1} .

171 *2.3. Cryogenic cell*

172 An environmental cell was designed and built at IPAG in order to simulate low CO_2 pressure and

173 low temperature (LP-LT) close to Martian atmospheric conditions. A heating resistance coupled
174 to a liquid N₂ circuit (77K) allows an efficient regulation of sample temperature from -180°C to
175 +100°C. Additionally, a turbomolecular vacuum pump and a CO₂ cylinder were connected to
176 reach a secondary vacuum and to inject a controlled CO₂ pressure into the reaction cell,
177 respectively. Figure 1 shows a schematic diagram of all main parts of the IR-CryoCell setup.

178 *2.4. Gas-solid carbonation experiments*

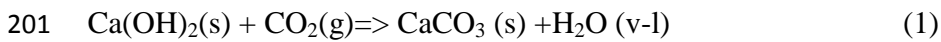
179 For these measurements, the reacting Ca(OH)₂ particles, stored at atmospheric conditions, were
180 manually deposited and compressed as a thin film on a KBr window. Then the KBr window was
181 carefully placed in the reaction cell to be assembled to the microscope. All carbonation
182 experiments have been carried out in presence of molecular water (adsorbed or crystallized as ice
183 depending on the carbonation temperature) which catalyze the carbonation process. The
184 carbonation temperatures used in this study were -10, 0, 10, 25 or 30°C and the CO₂ pressures
185 were typically 100, 1000 and 2000 mbar. This pressure is higher than Martian pressure to
186 accelerate the reaction due to a daily timescale limitation by the experimental setup. We note that
187 the CO₂ gas has been directly injected into the reaction cell in presence or absence of atmospheric
188 air. For the latter case, we started by fixing the water adsorbed onto the solid by cooling the cell
189 at -60°C before making a high vacuum pumping for 10 min in order to remove exclusively the air
190 from the reaction cell. After injection of CO₂ 10 to 15 infrared spectra have been collected as a
191 function of time until an apparent spectroscopic equilibrium state is reached (3-6h).
192 Complementary carbonation experiments have been carried out by using Mg hydroxide (brucite:
193 Mg(OH)₂) and the amorphous calcium silicate hydrate as solid reactants, but, for these cases the
194 carbonation temperature has been fixed at 25°C and 1 bar of CO₂ has been injected into the

195 reaction cell without air removal (more reacting system).

196 Each carbonation experiment has been repeated 2 times in order to verify its reproducibility. All
197 carbonation experiments and their physicochemical conditions are summarized in Table 1.

198 *2.5. Calculation of integrated band intensities*

199 The gas-solid carbonation of calcium hydroxide at low temperature (<30°C) in presence of
200 adsorbed water can be expressed by a global reaction as follows:



202 Generally, this global reaction is incomplete due to the formation of a protective carbonate layer
203 around the reacting particle which restricts or stops the CO₂ transfer at the grain or aggregate
204 scale (Montes-Hernandez et al., 2010a). In the present study, the integrated band intensities for
205 hydroxyl (-OH), carbonate (CO₃²⁻) and H₂O functional groups, concerning reaction (1) at an
206 instant *t* have been estimated by using a Trapezoidal rule integration. A wavenumber interval and
207 a characteristic continuum have been manually defined to determine the intensity of a given band
208 depending on the initial reactant. For example, in the gas-solid carbonation experiments with
209 Ca(OH)₂ particles, two continuums have been defined as linear segments over two different
210 spectral ranges, one for the -OH at 3640 cm⁻¹ and H₂O at 3450 cm⁻¹ band intensities and the other
211 for the H₂O at 1650 cm⁻¹ and CO₃²⁻ at 1420 cm⁻¹ band intensities (see Figure 2).

212

213 *2.6. Fitting of the kinetic experimental-calculated data for gas-solid carbonation*

214 Several kinetic models (first-order, pseudo-first-order, second-order, pseudo-second-order,
215 reversible one, irreversible one...) are generally used for fitting kinetic experimental data of
216 sorption and adsorption systems (Ho and McKay, 1999; Ho, 2006). For our experiments, we

217 have chosen pseudo-second-order model because it was successfully applied in previous studies
 218 (Montes-Hernandez and Geraud, 2004 ; Montes-Hernandez and Rihs, 2006; Montes-Hernandez
 219 et al., 2009, 2010a, 2010b, 2012a, 2012b) and can be adequately used to fit experimental data of
 220 carbonation process as demonstrated in Montes-Hernandez et al., 2009. This model reproduces a
 221 process consisting of a fast mass transfer followed by a second step of slower mass transfer until
 222 equilibrium is achieved. It can be written in its differential form as follows:

$$223 \quad \frac{dA^{CO_3}_{,t}}{dt} = Kc (A^{CO_3}_{,max} - A^{CO_3}_{,t})^2 \quad (2)$$

224 Where $A^{CO_3}_{,t}$ is the integrated band intensity for the carbonate group at a given time, t [minutes],
 225 corresponding to carbonation extent; $A^{CO_3}_{,max}$ is the maximum extend of carbonation at
 226 equilibrium; Kc is the rate constant of $Ca(OH)_2$ carbonation.

227 The second step (until equilibrium) is interpreted by as a passivation effect due to the formation
 228 of a protective carbonate layer (Montes-Hernandez et al., 2012a). In this study, the increase of
 229 integrated band intensity with time for the carbonate group (CO_3^{2-}), i.e. during gas-solid
 230 carbonation process, has been fitted by using a kinetic double-pseudo-second-order model. This
 231 model assumes two kinetic regimes due to the presence of two types of reactive surface sites. The
 232 integrated form of the double kinetic model is given by the following hyperbolic equation:

$$233 \quad A^{CO_3}_{,t} = \frac{(A^{CO_3}_{,max1})t}{(t_{1/2_1} + t)} + \frac{(A^{CO_3}_{,max2})t}{(t_{1/2_2} + t)} \quad (3)$$

234 Where $A^{CO_3}_{,t}$ is the integrated band intensity for the carbonate group at a given time, t [minutes],

235 corresponding to carbonation extent; $A^{\text{CO}_3}_{,\text{max}1}$ and $A^{\text{CO}_3}_{,\text{max}2}$ are the maximum extent of
236 carbonation at apparent equilibrium for both kinetic carbonation regimes, respectively; $t_{1/2,1}$ and
237 $t_{1/2,2}$ are the half-carbonation times for both kinetic carbonation regimes, respectively. In other
238 terms, the half-carbonation times represent the times after which half of the maximum of kinetic
239 carbonation regimes (expressed as maximum of integrated band intensities for carbonate group)
240 is obtained. The fitting of kinetics data allow an estimation of these parameters and was
241 performed by a non linear regression by least-squares method. These simple parameters are used
242 in this study to evaluate the kinetic effects of temperature, CO_2 pressure and nature of the solid
243 on the gas-solid carbonation process.

244 The activation energy (E_a , Table 1) of the reaction was calculated assuming an Arrhenius
245 behavior for the initial carbonation rate. We have used 4 points to calculate E_a for carbonation
246 experiments of portlandite (at the temperature of -10°C , 0°C , 10°C and 30°C , for both experiments
247 performed under 1 bar and 2 bar CO_2).

248

249 **3. Results**

250

251 *3.1. Gas-solid carbonation of $\text{Ca}(\text{OH})_2$ particles at low temperature ($<0^\circ\text{C}$)*

252 Very few experimental studies have characterized the carbonate formation or CO_2 mineralization
253 at the mineral-ice water- CO_2 interfaces on Earth and planetary cold-environments (e.g. Antarctic
254 and Mars surface). In our study, several gas-solid reactions carried out in the cryogenic cell
255 coupled to the infrared microscope reveal that carbonate formation or CO_2 mineralization is
256 possible at low temperature (-10°C and 0°C) using a simplified analogue $\text{Ca}(\text{OH})_2(\text{mineral})-$

257 water(adsorbed)-CO₂(gas) system (see Fig. 3 (c) to (f)). The results displayed in Figure 3 also
258 reveal that the carbonation extent, monitored *in-situ* by an increase of carbonate band intensity at
259 1420 cm⁻¹, is clearly inhibited by a decrease of temperature from 30°C to -10 °C. The increase of
260 integrated band intensity with time for the carbonate group at 1420 cm⁻¹ has been successfully
261 fitted by using the kinetic double-pseudo-second-order model. The experimental data and the
262 calculated fits for six experiments are plotted in Figure 4. This “a posteriori” modeling shows the
263 good fits of all the experimental data by such type of kinetic model (correlation coefficient, R
264 close to 1), and confirms the inhibition effect of temperature and the effect of relative humidity
265 on the carbonation extent and kinetic parameters (see also Table 1).

266 One additional carbonation experiment with Ca(OH)₂ particles was carried out at low CO₂
267 pressure (100 mbar) and at moderate temperature (25°C). For this case, the initial air contained
268 into the cell was previously removed by pumping to secondary vacuum at low temperature (-
269 60°C) as explained in the materials and methods section. Here, a significant carbonation is
270 observed after 4 minutes of Ca(OH)₂-CO₂ interaction followed by a very slow carbonation step
271 until an apparent spectroscopic equilibrium is possibly reached (about 6h) (see Fig. 5). These
272 experimental data have been also successfully fitted by using the kinetic double-pseudo-second-
273 order model. A last carbonation experiment was performed at a CO₂ pressure of 2 bar (at 25 °C)
274 in order to compare with the low CO₂ pressure experiments. A significant carbonation was
275 observed during all the experiment (see Fig. 6), which was fitted with the kinetic-pseudo-second
276 order model.

277 3.2. Gas-solid carbonation of Mg hydroxide

278 The gas-solid carbonation depends also on the nature of the solid. For this reason one other

279 powdered material, Mg hydroxide (synthetic brucite), was investigated specifically at higher
280 reactive conditions (25°C and 1bar of CO₂, in presence of air). To form Mg carbonates, the most
281 simple materials as starting reactant are binary oxides or hydroxides in the precursor material.
282 Brucite particles are found to be only slightly carbonated at these T-P_{CO₂} conditions after 5.5h of
283 Mg(OH)₂-CO₂ interaction (see Figure 7). These *in-situ* infrared measurements clearly reveal that
284 the Mg hydroxide (brucite) is more chemically stable than Ca hydroxide (portlandite) under a
285 CO₂-rich atmosphere at a given relative humidity. In summary, the gas-solid carbonation of Ca
286 and Mg hydroxides depends on the experimental conditions employed (i.e. T, P_{CO₂}, relative
287 humidity) and on the intrinsic properties of solid (i.e. hydrophilicity, particle size, specific surface
288 area, and chemical stability).

289 Finally, a kinetic regime and the maximum carbonation extent at an apparent equilibrium
290 ($A^{\text{CO}_3}_{,\text{max}1} + A^{\text{CO}_3}_{,\text{max}2}$) is successfully determined by using a kinetic double-pseudo-second-order
291 model (see Fig. 7 (c)).

292 3.3. Gas-solid carbonation of an amorphous calcium silicate hydrate

293 A last materials, an amorphous calcium silicate hydrate, has been investigated at higher reactive
294 conditions (25°C and 1bar of CO₂, in presence of air) to test the gas-solid carbonation efficiency.
295 The amorphous calcium silicate hydrate, is significantly carbonated via gas-solid carbonation at
296 the above mentioned T-P_{CO₂} conditions after 8h of reaction (see Figure 8), which suggests
297 chemical stability has a significant impact on the efficiency of the carbonation.

298 Finally, a kinetic regime and the maximum carbonation extent at apparent equilibrium ($A^{\text{CO}_3}_{,\text{max}1}$
299 $+ A^{\text{CO}_3}_{,\text{max}2}$) is also successfully determined by using the kinetic double-pseudo-second-order
300 model (see Fig. 8 (c)).

301

302 **4. The mechanism of carbonation**

303 All the experiments with the Ca and Mg hydroxides show an increase of the band intensities of
304 carbonates, at low temperature and low pressure. In this study we assume that part of the water
305 initially adsorbed onto Ca(OH)₂ particles was partially crystallized by cooling when the
306 temperature is negative (<0°C). The presence of an ice layer limits the access of CO₂ molecules
307 to nanopores, and therefore limiting the CO₂ access to the local CO₃²⁻ production
308 (CO₂(g)+H₂O(adsorbed)=>CO₃²⁻+2H⁺) required to form a carbonate layer around the Ca(OH)₂
309 particles (see also: Montes-Hernandez et al. 2010a). Strictly speaking, the relative humidity is not
310 controlled in our experiments; however, two experiment protocols implying atmospheric vapor
311 have been designed, firstly, direct injection of CO₂ gas into the reaction cell initially filled with
312 air, i.e. at lab relative humidity (CO₂-air system) and secondly, the injection of CO₂ gas after
313 removal of the air by secondary vacuum pumping at low temperature (-60°C) (CO₂ system). The
314 difference between these experiments could explain why the carbonation extent decreases when
315 the initial air (contained into the reaction cell) is removed (see comparisons (c) and (d) or (e) and
316 (f) in Fig. 3). We can assume a similar relative humidity of the lab room for all experiments. The
317 relative humidity has clearly an impact on the carbonation efficiency, the experiments without air
318 (very low relative humidity) showing a lower amount of carbonation at low temperature.

319 The fit of the data by the kinetic model assumes two kinetic regimes, usually due to the presence
320 of two types of reactive surface sites. In our carbonation experiments, the formation of a hydrated
321 carbonate layer around the core of reacting Ca(OH)₂ particles produces a complex passivation
322 step, possibly perturbed by three simultaneous physicochemical processes: (1) solid state
323 transformation from hydrated calcium carbonate to calcite and/or from aragonite to calcite, (2)
324 partial expelling of produced molecular water during the carbonation process (see Eq. (1)) and (3)

325 local acidification by an excess of molecular water in pores or onto surfaces ($\text{H}_2\text{O}(\text{produced}) +$
326 $\text{CO}_2(\text{g}) \Rightarrow \text{HCO}_3^- + \text{H}^+$). In summary, the complex kinetic behavior related to gas-solid
327 carbonation of $\text{Ca}(\text{OH})_2$ particles is successfully described applying two kinetic regimes. A
328 schematic representation of this carbonation process is illustrated in Figure 9. The rate of
329 carbonation depends on the access to the nanopores of the material by the CO_2 . These pores have
330 to be water-unsaturated to facilitate access of the CO_2 gas to react with the minerals. The pressure
331 has a strong impact on the rate and yield of carbonation. In the case of the low pressure
332 experiments, a two stage kinetic model was shown to fit the data. Experiments revealed a fast
333 carbonation during a short time (stage 1) followed by a slower carbonation (stage 2). The
334 magnitude of carbonate formation is high in stage 1 and lower in stage 2. In the case of the
335 experiments performed at higher CO_2 pressure (2 bar) (fig. 6.c) a two stage reaction is also
336 observed. However, unlike the low pressure experiments, the magnitude of carbonation achieved
337 in stage 2 is quite significant.

338 For the low pressure experiments, we suspect that the intraparticle diffusion of CO_2 , possibly
339 limited by the low gas pressure in the system (100 mbar of CO_2), is the rate limiting step due to
340 the carbonate layer which strongly reduce the diffusion of the gas. This rate limiting step is no
341 more observed at high CO_2 pressure (>20 bar). In this case, the $\text{Ca}(\text{OH})_2$ particles are completely
342 carbonated, leading to the formation of calcite nano-crystals (Montes-Hernandez et al., 2010b).
343 We can assume a correlation between the pressure and the thickness of the layer that transforms
344 to carbonate by gas-solid reaction. The effect of CO_2 pressure observed in explained by the
345 presence of passivation step and the formation of carbonate layer through which CO_2 molecules
346 have to diffuse. Therefore in the case of an uncarbonated material, the effect of CO_2 pressure on
347 the initial reaction rate is expected to be moderate. Although the CO_2 pressure on Mars (about 10

348 mbar) is lower than CO₂ pressures used in our experiments (100 mbar), it is likely that our results
349 can be extrapolated to Martian atmospheric CO₂ pressure.

350

351 Unfortunately, the gas-solid carbonation mechanism of amorphous calcium silicate hydrate is not
352 elucidated due to its unknown atomic organization. However, we assume that the abundant
353 molecular water adsorbed onto the solid plays a crucial role to start the gas-solid carbonation
354 process at the investigated conditions. The *in-situ* infrared measurements reveal two important
355 insights: (1) The expelling of pre-existent molecular water in/on the solid towards the gas phase
356 during the carbonation process. This is attested by a clear decrease of the stretching and bending
357 band intensities of water (see Fig. 8 (a) and (b)), (2) Similar to carbonation of Ca hydroxide, the
358 formation of calcite and aragonite are mainly identified, the formation of hydrated calcium
359 carbonate being only suspected (see also Montes-Hernandez et al. 2010a)

360

361

362 **5. Discussion**

363

364 Carbonates have been found on Mars in two kinds of geological settings: (i) outcrops of
365 carbonates, identified in the Nili Fossae region (Ehlmann et al., 2008), in the central peak of
366 Leighton crater (Michalski and Niles, 2010) and in the Columbia Hills of Gusev crater (Morris et
367 al., 2010); and (ii) carbonates-bearing dust, identified by the TES instrument (Bandfield et al.,
368 2003) and the phoenix lander (Boynton et al., 2009). In the case of the outcrops from the
369 Columbia Hills and Nili Fossae, carbonates are present as major components (16 to 34 wt % in
370 the case of the Columbia Hills, about 80 % in the case of Nili Fossae), and their derived

371 chemistry is similar to that of carbonates found in Martian meteorites (Mittlefehldt, 1994), i.e.
372 Fe-Mg carbonates. The association of these carbonate outcrops with phyllosilicates advocate for a
373 possible hydrothermal origin of these carbonates, a phenomenon that has been reproduced in
374 laboratory experiments (Golden et al., 2000) and that is observed in some terrestrial hydrothermal
375 systems (Treiman et al., 2002; Brown et al., 2010). However, it is well known that terrestrial
376 alteration of mafic rocks can produce brucite as a primary alteration product (Xiong and Snider
377 Lord, 2008), which should readily transform into carbonate by interaction with the Martian
378 atmosphere, according to our experiments. The observed carbonates outcrops could rather be
379 former outcrops of brucite-rich sedimentary rocks, that were subsequently altered to carbonates
380 by interaction with the atmosphere.

381 In the case of carbonates observed in the Martian dust, both magnesite (Bandfield et al., 2003)
382 and calcite (Boynton et al., 2009) have been reported, and their typical abundance is below 5 %.
383 Although aqueous formation has received widespread attention for this type of occurrence of
384 carbonates on Mars, we propose gas-solid reaction as a possible formation mechanism. Calcite
385 formation at the dust-CO₂ interfaces requires a source of calcium (e.g. Ca binary oxides or an
386 amorphous metastable Ca silicate) possibly coming from volcanic activity (Shaheen et al., 2010),
387 mechanical erosion or extra-Martian particulate matter (including meteorite impacts, interstellar
388 dusts). A large diversity of phyllosilicates and hydrated phyllosilicates was found on the Martian
389 surface (Mustard et al., 2008; Jänchen et al., 2006; Fairén et al., 2009; Murchie et al., 2009;
390 Ehlmann et al., 2011). As we have shown, the presence of molecular water is also required
391 because hydration of the Ca precursor is assumed to be a crucial step prior to the carbonation
392 process. Laboratory studies of Martian analogs suggest that adsorbed water should be present in
393 significant amount within the Martian soil (Pommerol et al., 2009; Beck et al., 2010; Jänchen et

394 al., 2006) and adsorbed water has been also detected by infrared spectroscopy (Poulet et al.,
395 2009). In addition, the gamma rays and neutrons spectrometers on Mars Odyssey have shown
396 evidence for the presence of water in the first meter of the martian subsurface (Feldman et al.,
397 2004). A simplified scenario for calcite formation at the dust-CO₂ interfaces and its natural
398 deposition on the soil is schematically illustrated in Fig. 10. In this scenario we assume that the
399 precursor, a calcium hydroxide with adsorbed water, is produced by atmospheric alteration of
400 volcanic CaO particles in the atmosphere. Reactant minerals such as portlandite could be difficult
401 to detect on Mars by reflectance spectroscopy due to the carbonate layer around the calcium
402 hydroxide. In the case of hydromagnesite, the presence of brucite is required somewhere on
403 Mars, which would be subsequently transformed to carbonates, eroded, and transported. As we
404 stated earlier, brucite should form in association with phyllosilicates during the aqueous alteration
405 of mafic rocks.

406 The efficiency of carbonates synthesis on Mars by gas-solid reaction will depend on the
407 mineral substrate (as we showed, brucite, portlandite and larnite have distinct synthesis kinetics),
408 the local temperature, the atmospheric humidity, and likely the atmospheric pressure (which can
409 substantially vary with season as well as with topography). Even at temperatures below the frost
410 point, carbonates synthesis can occur by gas-solid reaction, on a daily timescale (table 1). On
411 Mars, the water vapor pressure is low ($P_{\text{H}_2\text{O}}$ about 10 Pa) and the frost point is depressed with
412 regard to that on Earth. Given the present knowledge of the water vapor surface pressure, a
413 typical value of 200 K is found for the frost point on the Martian surface (Schorghofer and
414 Aharonson., 2005). Such temperature typically corresponds to seasonal average around 60° in
415 latitude.

416 The relative humidity on Mars fluctuates on a daily basis. At the Phoenix landing site (Smith P.H.

417 et al., 2009) it is measured around 5% during Martian day time, close to saturation early night
418 and saturating at the end of the night. Carbonate synthesis will be accelerated by a high
419 atmospheric humidity, which can occur during the warmer season. TES instrument on Mars
420 Global Surveyor has water vapor evolution during 2 martian years, and the maximum was found
421 during midsummer in the northern hemisphere with 100 pr- μm (Smith M.D., 2002,2004). The
422 maximum water vapor measured by Mars Express instruments (OMEGA and SPICAM) is found
423 during midsummer, around 60 pr- μm content. This maximum is observed at latitude 75-80°N
424 and longitudes 210-240°E (Fouchet et al., 2007; Fedorova et al., 2006; Melchiorri et al.,2007)),
425 an area is close to Phoenix landing site (latitude 68°N and longitude 233°E).

426 Current Global Circulation Models (GCM) of Mars can be used to determine the optimal
427 locations and times for the gas-solid synthesis of carbonates. Simulations with Mars Climate
428 Database (Forget et al., 1999, 2006) estimated high relative humidity (around 70%) and
429 temperatures close to -20°C during mid summer ($L_s = 120^\circ$) in the Phoenix landing site area (at 12
430 a.m). These conditions are sufficient to initiate the carbonation reaction according to our
431 experiments, and carbonates were observed in the Phoenix soil (Smith et al., 2009)

432

433 Mars is not the only extra-terrestrial body where carbonates have been detected, this is
434 also the case of Ceres, the largest asteroid in the main belt. Its shape is close to hydrostatic
435 equilibrium and its bulk density suggests the presence of ice in its interior (Thomas et al., 2005).
436 The surface of Ceres shows a well-resolved 3- μm absorption band, which interpretation has been
437 debated (Lebofsky et al., 1981; Vernazza et al., 2005; Rivkin et al., 2006; Rivkin et al., 2011;
438 Beck et al., 2011). In a recent study, Milliken and Rivkin (2009) combined NIR and MIR
439 observations of Ceres' surface and successfully modeled both spectral regions with a combination

440 of brucite, carbonate and a Fe-rich phyllosilicate. Such a mineralogical assemblage was explained
441 by aqueous alteration of mafic silicates in the presence of CO₂, by analogy with the processes
442 inferred from the mineralogy of hydrated meteorites that can present a significant amount of
443 carbonates (Zolensky et al., 2002). If brucite is rare in the mineralogy of hydrated chondrites, it is
444 a common product of aqueous alteration of terrestrial rocks. The condition of brucite formation is
445 specific in terms of T, pH and pO₂, and source rock. If the formation of carbonate by reaction of
446 brucite with water is possible, gas-solid reaction cannot be excluded. This mechanism could
447 occur at some depth in the asteroid body, where CO₂ pressure can build-up. However, because of
448 the low temperature at the surface of Ceres, long timescales are expected for such a process.
449 Further consideration would require an accurate knowledge of the kinetics of the gas-solid
450 carbonation.

451

452 Finally, the carbonate synthesis mechanism that we described is certainly active on Earth,
453 where carbonate minerals played an important role in the planet evolution. Many studies are
454 available about carbonate reactivity and synthesis in liquid-water but information on its behavior
455 at sub-zero temperatures (for example solubility in frozen water) are sparse. The results we
456 obtained reveal that gas-solid carbonation can occur at sub-zero temperature, in the presence of
457 gaseous CO₂ and H₂O. These conditions are present on Earth in arctic regions and in the upper
458 atmosphere. This mechanism can thus occur on the availability of the adequate precursor.

459 The presence of oxygen isotope anomalies in carbonates from terrestrial aerosols
460 (Shaheen et al., 2010) suggests a carbonation by exchange with ozone. Such a result suggests a
461 possible formation of carbonate by chemical reaction in the upper atmosphere, from a CaO
462 precursor. We can propose hydration of CaO_(s) by H₂O_(g), and successive reaction of Ca(OH)_{2(s)}

463 with $\text{CO}_{2(\text{g})}$ as a formation mechanism of these carbonates.

464 Calcium carbonate and carbonate hydrates have been found in arctic ice (Dieckmann et
465 al., 2008; Sala et al. 2008). The formation mechanism of these carbonates is a matter of active
466 research, since it could provide a major CO_2 sequestration process. Hydrous carbonates (for
467 instance ikaite) have been proposed to originate by precipitation during sea-ice formation, as
468 suggested by thermodynamical calculations. Anhydrous carbonates can have an origin as primary
469 aerosols, with a synthesis mechanism possibly similar to the one described in the previous
470 paragraph. In addition, in situ gas-solid formation is possible, depending on the availability of an
471 adequate precursor (as we showed here, an Ca or Mg hydroxides, or Ca-rich amorphous
472 silicates).

473

474 **6. Conclusion**

475 In this study, we designed an original experimental method to form carbonates via gas-solid
476 reaction in presence of adsorbed water. We used an infrared microscope coupled to a cryogenic
477 reaction cell (IR-CryoCell setup) to investigate this process with 3 different carbonate precursors
478 (Ca hydroxide (portlandite), Mg hydroxide (brucite), and an amorphous calcium silicate hydrate).
479 We demonstrated for the first time that calcium carbonate can be formed at low temperature
480 ($<0^\circ\text{C}$) via gas solid carbonation of Ca hydroxide. Both amorphous Ca silicate hydrate and Ca
481 hydroxide were significantly carbonated at the investigated T- P_{CO_2} conditions. Conversely, only a
482 very slight gas-solid carbonation of Mg hydroxide particles was detected by IR spectroscopy. We
483 extracted the kinetic parameters of the reaction from our measured carbonation curves, following
484 a kinetic double-pseudo-second-order model. From these results we can clearly state that the
485 conditions for gas-solid carbonation exist on Mars, and that this process could be the source of

486 the detected Ca and Mg carbonates found in the Martian dust and soil. These carbonates can be
487 synthesized from a brucite precursor (a common hydrothermal product), from volcanic derived
488 aerosols, as well as from extraterrestrial dust. This mechanism should be considered in future
489 global modeling of the carbon cycle of the red planet, and might also be active in cold terrestrial
490 deserts.

491

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496

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703 Table 1. Summary of the experiments with their experimental conditions and the corresponding
 704 kinetic parameters determined for gas-solid carbonation.

Exp.	Starting material	Gas pressure	temperature	$A^{\text{CO}_3}_{\text{max1}}$ (a.u.)	$A^{\text{CO}_3}_{\text{max2}}$ (a.u.)	$t_{1/2_1}$ (minutes)	$t_{1/2_2}$ (minutes)	Ea (kJ/mol)
1	portlandite	2 bar CO ₂	-10°C	1.8	4.6	19.6	599.6	43
2	portlandite	2 bar CO ₂	0°C	6.4	0.8	8.9	8.9	
3	portlandite	2 bar CO ₂	10°C	5.3	11.4	61.8	61.8	
4	portlandite	2 bar CO ₂	25°C	29.3	54.8	2.5	126.6	
5	portlandite	1 bar CO ₂ + air	-10°C	2.5	5	3.7	33622	75
6	portlandite	1 bar CO ₂ + air	0°C	5.4	10.2	6.3	28.5	
7	portlandite	1 bar CO ₂ + air	10°C	15.8	3.1	3.5	180.1	
8	portlandite	1 bar CO ₂ + air	30°C	26	21.2	3.5	5.8	
9	portlandite	100 mbar CO ₂	25°C	19.8	10.9	0.8	419.1	
10	brucite	1 bar CO ₂ + air	25°C	9.5	4.8	3	184.1	
11	Amorphous Ca silicate hydrate	1 bar CO ₂ + air	25°C	8.9	68	<0.5	13.6	

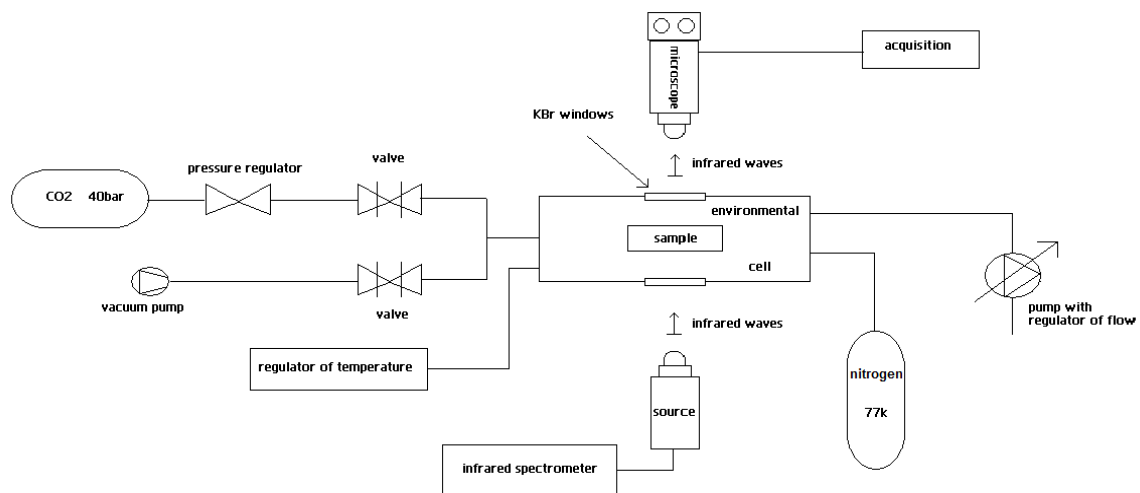
Ea was calculated with Arrhenius equation. For experiments with 2 bar CO₂, we exclude the point of 10°C due to his incoherence with Arrhenius equation.

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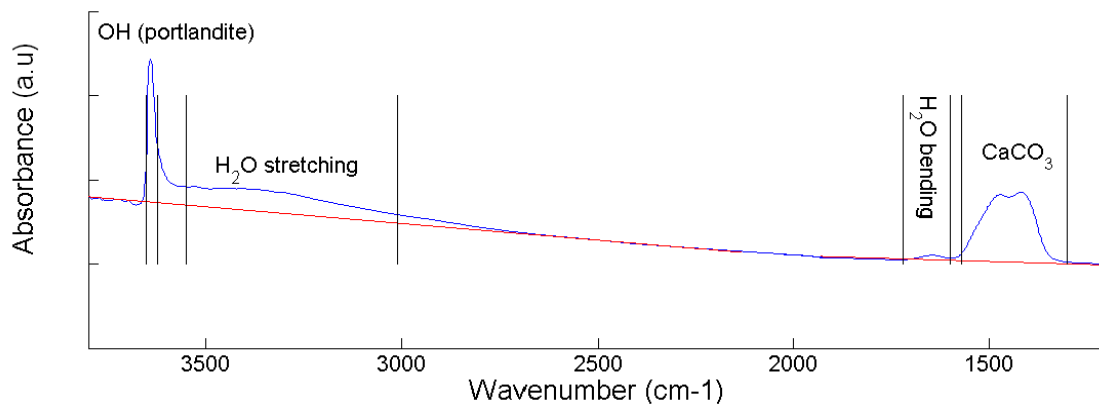
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714 **Figure 1** . Schematic representation of the IR-CryoCell experimental setup, showing the main
 715 parts such as temperature regulator, environmental cell, infrared microscope, valves, vacuum
 716 pump, CO₂ cylinder, liquid N₂ reservoir.

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722 **Figure 2** . Schematic representation for the calculation of the integrated band intensities of each
 723 functional group (-OH, H₂O, MCO₃), showing the continuum (in red) on an IR spectrum of
 724 portlandite.

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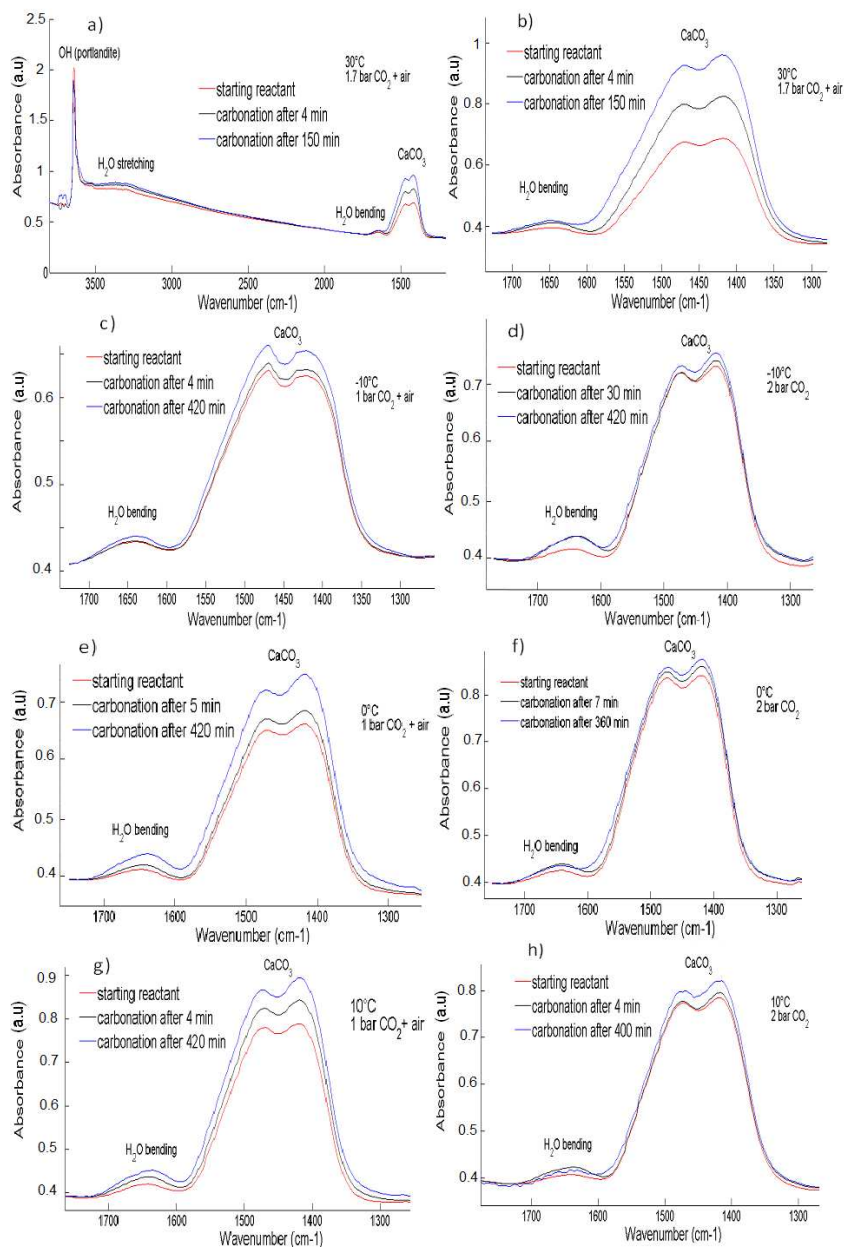
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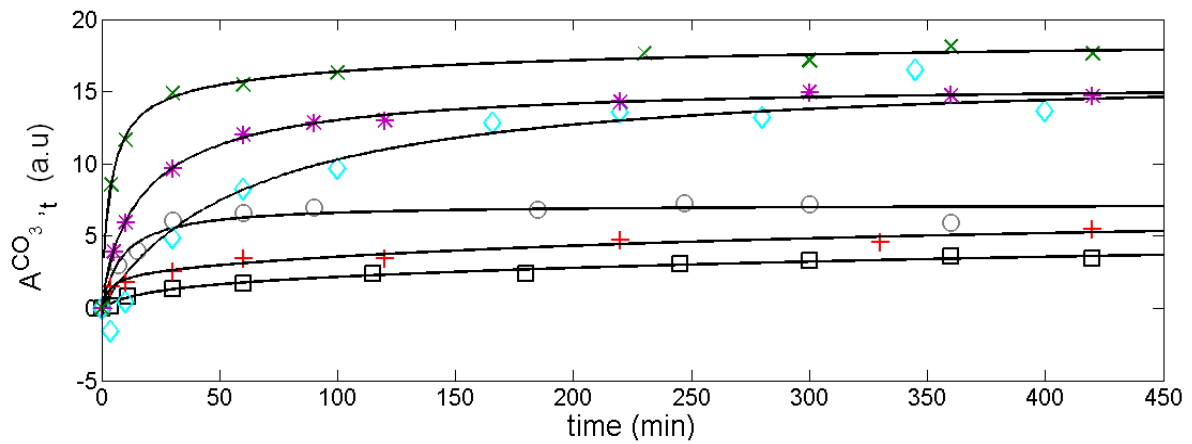
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739 **Figure 3.** Evolution with time of the IR spectrum of calcium carbonate during carbonation at
 740 different temperatures and CO₂ pressures. a) Full IR spectrum of portlandite at 30°C with 1.7 bar
 741 of CO₂ in presence of air; (b) Carbonate band at 30°C under 1,7 bar of CO₂ with air; (c) at -10°C
 742 under 1bar of CO₂ with air ; (d) at -10°C under 2 bars of CO₂; (e) at 0°C under 1bar of CO₂ with
 743 air ; (f) at 0°C under 2bars of CO₂; (g) at 10°C under 1bar of CO₂ with air; (h) at 10°C under
 744 2bars of CO₂.



- ◇ 10°C, 2bar CO2
- 0°C, 2bar CO2
- -10°C, 2bar CO2
- × 10°C, 1bar CO2 + air
- * 0°C, 1bar CO2 + air
- + -10°C, 1bar CO2 + air
- Kinetic double-second-order-model

Kinetic double-second-order-model

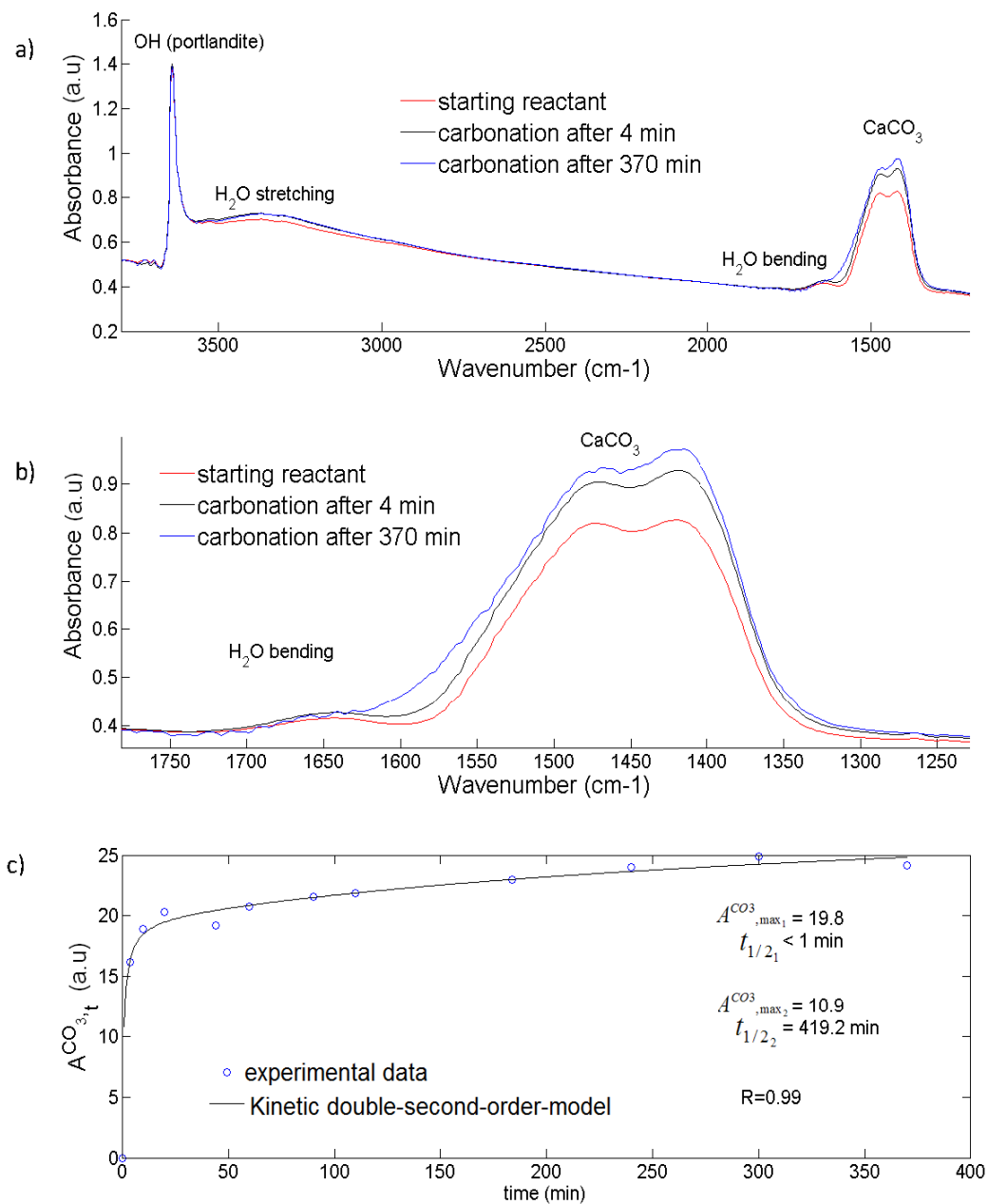
$$A^{CO_3}_t = \frac{(A^{CO_3}_{,max_1})t}{t_{1/2_1} + t} + \frac{(A^{CO_3}_{,max_2})t}{t_{1/2_2} + t}$$

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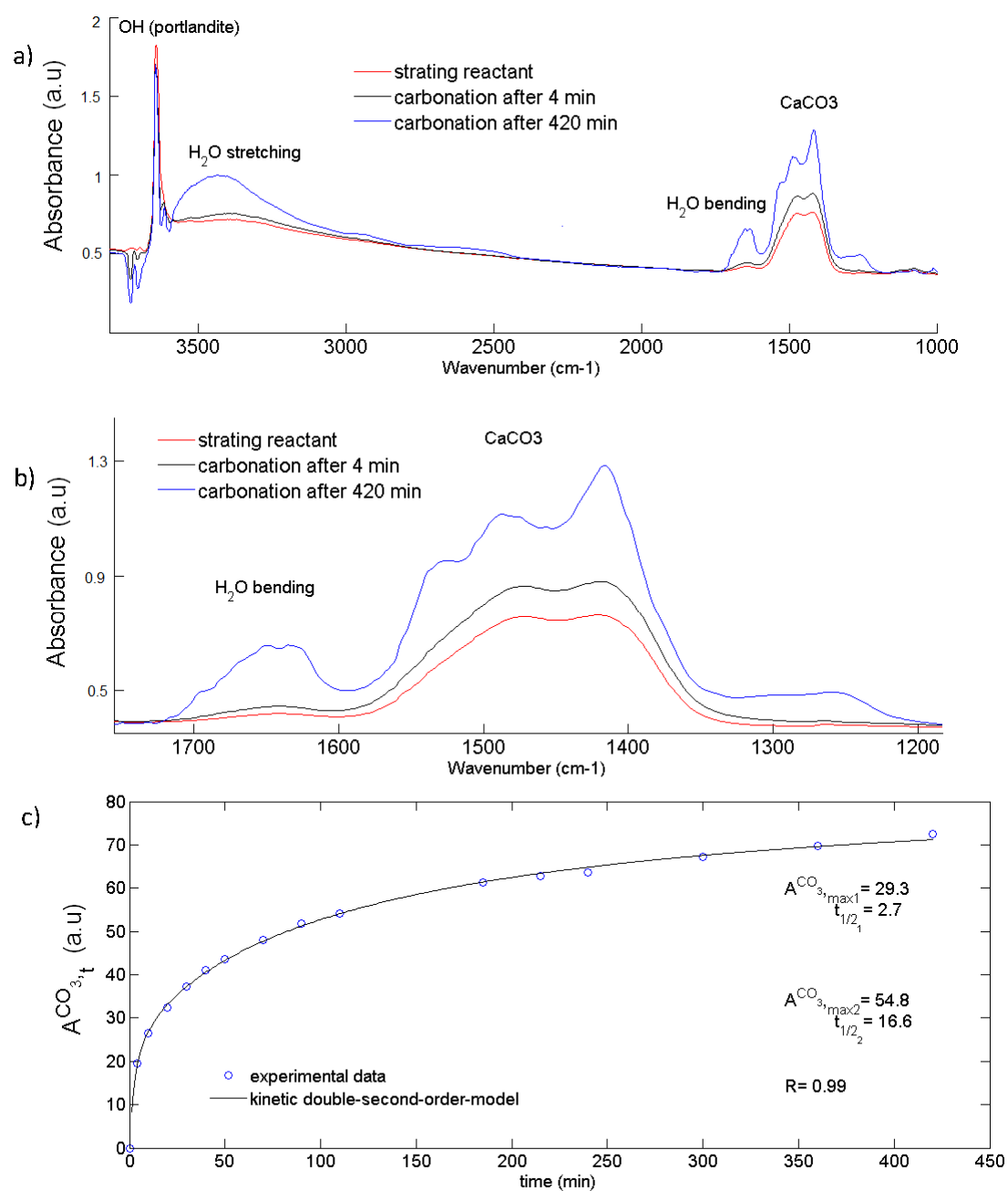
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748 **Figure 4.** Fits of the experimental kinetic data (carbonate band intensity) for gas-solid
 749 carbonation of Ca hydroxide (portlandite) in various experimental conditions by using a kinetic
 750 double-pseudo-second-order model and applying the non-linear least squares method.

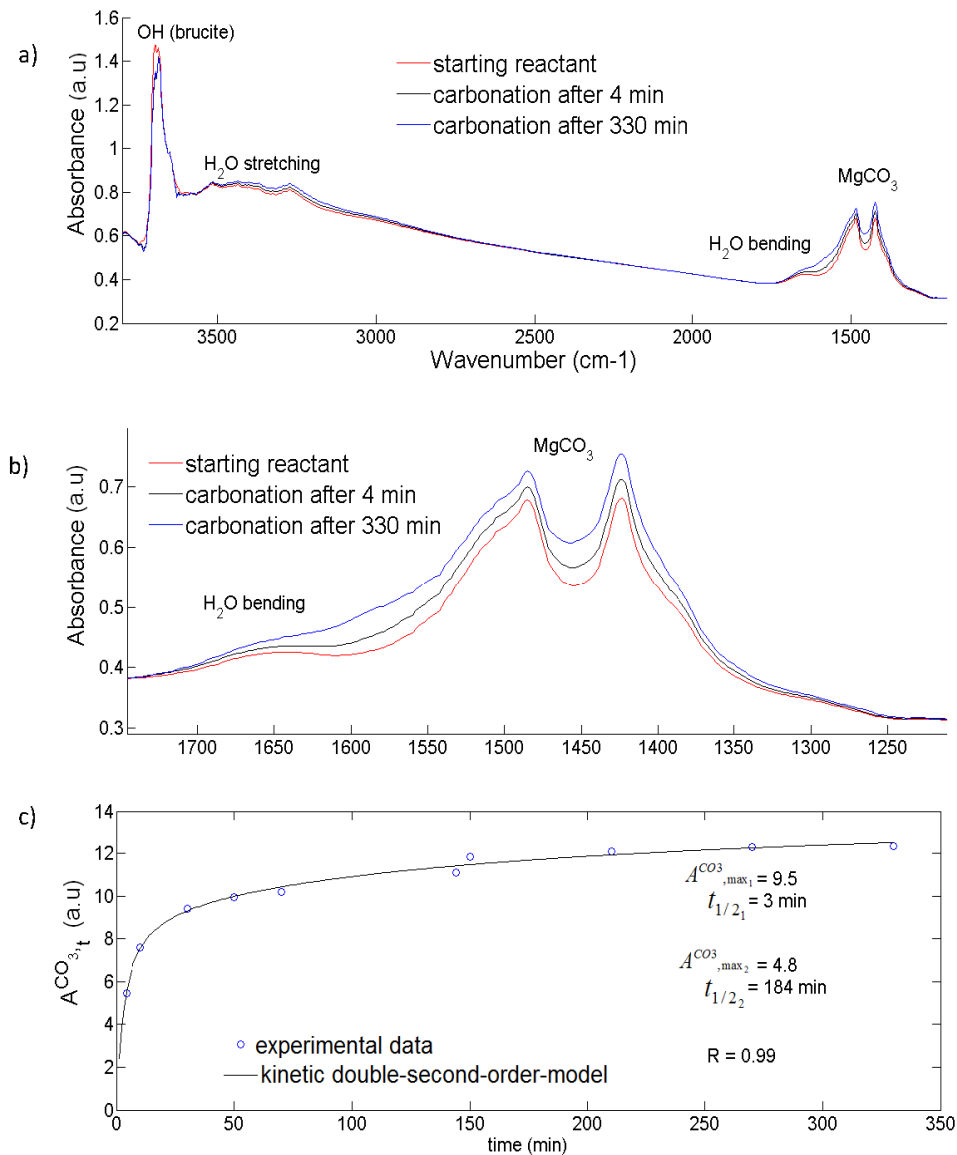


751 **Figure 5.** Evolution with time of the IR spectrum of Ca hydroxide (portlandite) during
 752 carbonation at 25°C under 100mbar of CO₂: a) Full spectrum. b) Band of the carbonate group. c)
 753 Fit of the experimental kinetic data (carbonate band intensity) for gas-solid carbonation by using
 754 a kinetic double-pseudo-second-order model and applying the non-linear least squares method.



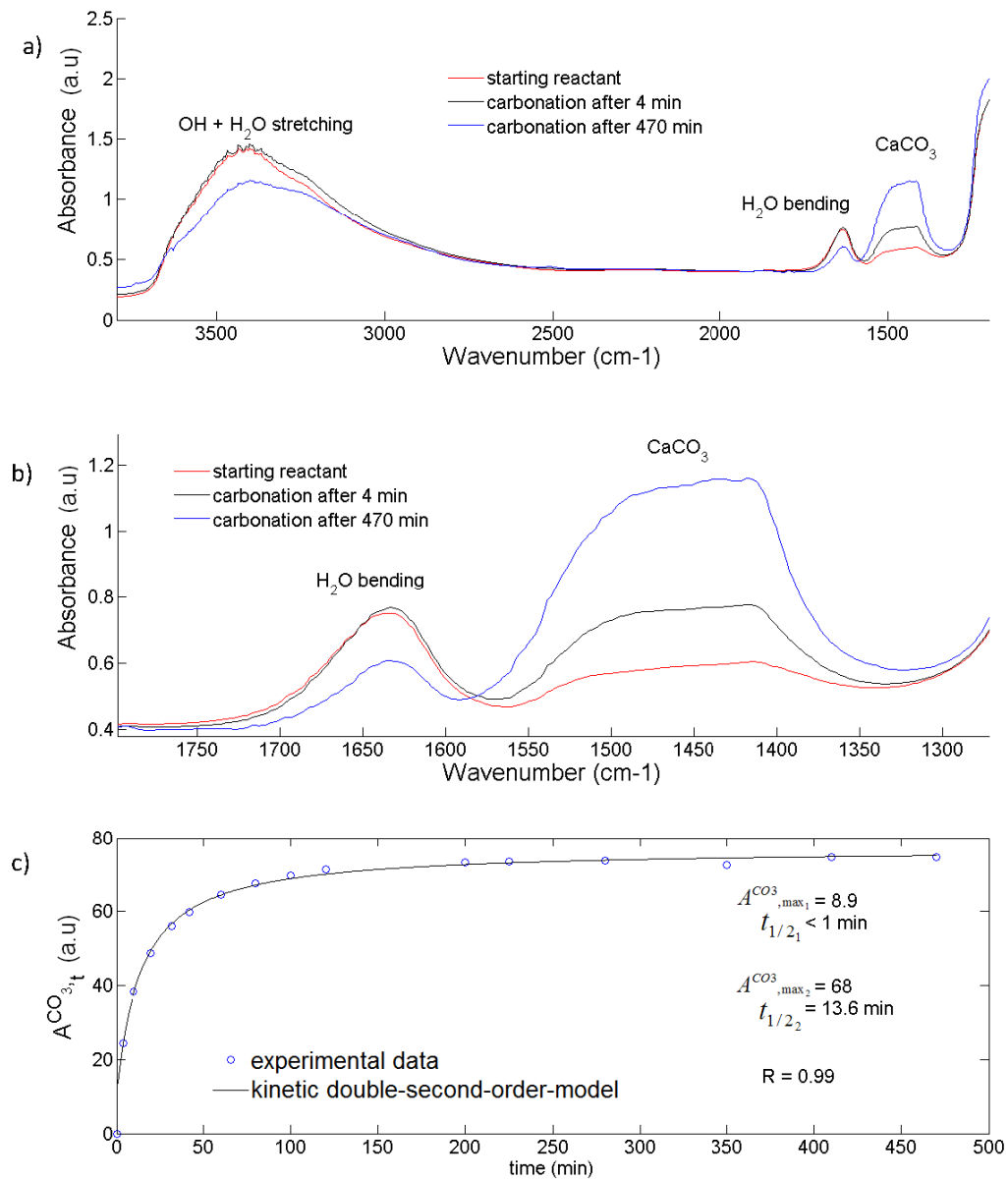
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756 **Figure 6.** Evolution with time of the IR spectrum of Ca hydroxide (portlandite) during
 757 carbonation at 25°C under 2 bar of CO₂: a) Full spectrum. b) Band of the carbonate group. c) Fit
 758 of the experimental kinetic data (carbonate band intensity) for gas-solid carbonation by using a
 759 kinetic double-pseudo-second-order model and applying the non-linear least squares method.



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761 **Figure 7.** Evolution with time of the IR spectrum of the Mg hydroxide (brucite) during
 762 carbonation at 25°C under 1bar of CO₂ with air: a) Full spectrum. b) Band of the carbonate group
 763 c) Fit of the experimental kinetic data (carbonate band intensity) for gas-solid carbonation by
 764 using a kinetic double-pseudo-second-order model and applying the non-linear least squares
 765 method.



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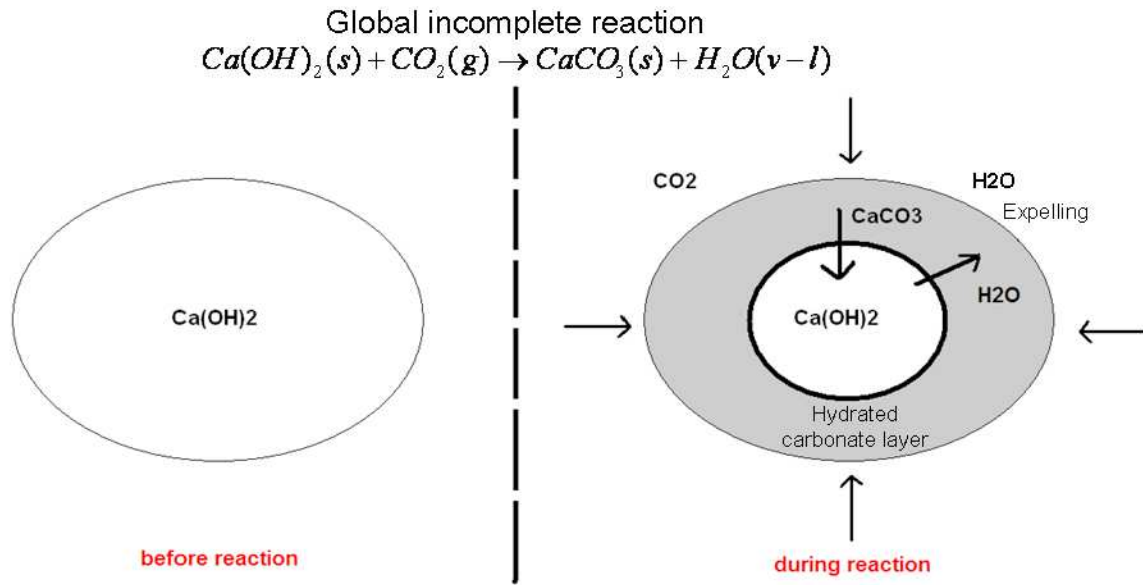
767 **Figure 8.** Evolution with time of the IR spectrum of amorphous Ca silicate hydrate during
 768 carbonation at 25°C under 1bar of CO₂ with air: a) Full spectrum. b) Band of carbonate group. c)
 769 Fit of the experimental kinetic data (carbonate band intensity) for gas-solid carbonation by using
 770 a kinetic double-pseudo-second-order model and applying the non-linear least squares method.

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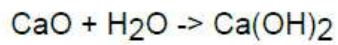
776 **Figure 9.** Schematic representation of the gas-solid carbonation of Ca hydroxide, showing the
777 growth of a hydrated calcium carbonate layer and the expelling of molecular water.

Binary oxides or amorphous silicates

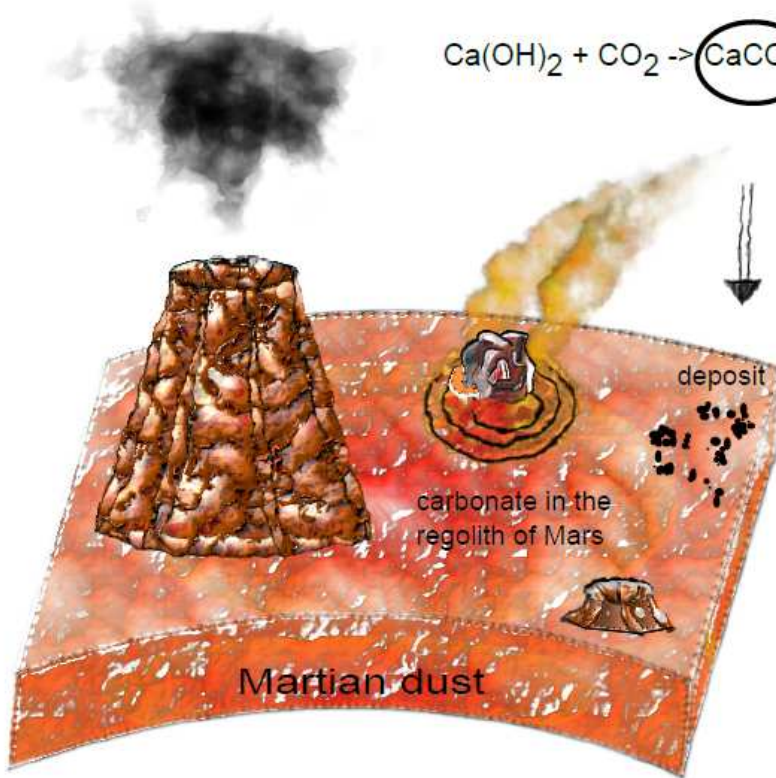
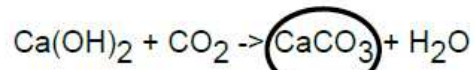


atmospheric conditions are enough to start the carbonation process

spontaneous hydration of precursor



unstable and react with atmospheric CO₂



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781 **Figure 10.** A schematic representation of a possible current formation mechanism at dust-CO₂
782 interfaces of the calcium carbonate found at the Martian surface.