

# The APSIS EXPERIMENT: Simulating the VUV Photochemistry of the Upper Atmosphere of Titan

Sarah Tigrine, Nathalie Carrasco, Ludovic Vettier, Olivia Chitarra, Guy Cernogora

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Sarah Tigrine, Nathalie Carrasco, Ludovic Vettier, Olivia Chitarra, Guy Cernogora. The APSIS EXPERIMENT: Simulating the VUV Photochemistry of the Upper Atmosphere of Titan. Titan Aeronomy and Climate (TAC) Workshop, Jun 2016, Reims, France. . insu-01348915

### HAL Id: insu-01348915 https://insu.hal.science/insu-01348915

Submitted on 26 Jul 2016

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## The APSIS EXPERIMENT: Simulating the VUV Photochemistry of the Upper Atmosphere of Titan

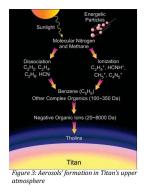
S. Tigrine<sup>1,2</sup>, N. Carrasco<sup>1</sup>, L. Vettier<sup>1</sup>, O. Chitarra<sup>1</sup>, G. Cernogora<sup>1</sup>

<sup>1</sup>LATMOS, Université Versailles St. Quentin, UPMC Univ. Paris 06, CNRS, 11 Bvd. d'Alembert, 78280 Guyancourt, France <sup>2</sup>Synchrotron SOLEIL, l'Orme des Merisiers, St Aubin, BP 48, 91192 Gif sur Yvette Cedex, France



Figure 1 & 2: Titan Cassini's imager





Titan, the largest moon of Saturn, has a dense atmosphere whose upper layers are mainly composed of methane (CH<sub>4</sub>) and molecular nitrogen (N<sub>2</sub>). The Cassini mission revealed that the interaction between those molecules and the solar VUV radiation, as well as the electrons from Saturn's magnetosphere, leads to a complex chemistry above an altitude of 800km [1].

This naturally ionized environment contains heavy organic molecules like benzene (C<sub>6</sub>H<sub>6</sub>) even at altitudes higher than 900 km [2]. This is consistent with an initiation of the aerosols in Titan's upper atmosphere. Moreover, some N-bearing molecules of pre-biotic interest such as NH3 have been detected by the instruments; but in quantities that do not match the theoretical models [3] .

The presence of those molecules makes Titan a natural laboratory to witness and understand prebiotic-like chemistry but despite all the data collected, all the possible chemical processes in such a hydrocarbon-nitrogen-rich environment are not precisely understood.

→ High interest of Titan's ionospheric chemistry experiments with pure-photochemistry simulations

#### **The APSIS Experiment**

Atmospheric Photochemistry SImulated by Surfatron

In order to reproduce the efficient photochemistry occurring in this kind of upper atmospheres, we designed a gas reactor named APSIS. This reactor is to be coupled with a VUV photon source (figure 5) as N, needs wavelengths shorter than 100 nm in order to be

The APSIS reactor is described in details in [4]. Briefly, it is a stainless steel chamber of dimensions 500 mm x 114 mm x 92 mm where the reactive 95:5 N2:CH4 mixture is introduced via a gas inlet (figure 4)

#### The VUV Source

At the LATMOS laboratory, a surfatron source has been developed using noble gases in a microwave discharge. For example, neon has two resonance lines at 73.5 and 74.3 nm which allow us to dissociate and/or ionize both CH4 and N2.

The photon flux of the source depends on the neon pressure and the micro-wave power delivered. We calculated it for the 73.5-nm neon line: it varies between a maximum of 4×1014 ph.s-1.cm-2 (1.7 mbar and 100W) and a minimum of 2×1013 ph.s-1.cm-2 (0.4 mbar and 40W)

led window-less with the VUV source

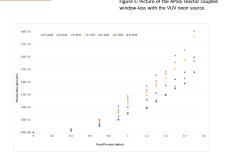




Figure 6: Photon flux of the VUV source versus the neon pressure and the micro

#### RESULTS

P<sub>APSIS</sub>= 0.5 mbar

\_\_= 1 mbar

P<sub>w</sub>=80W

Diagnosis by mass spectrometry: HIDEN HPR-20 QIC mass spectrometer

Figure 7 shows analogue mass spectra prior (black curve) and after (red curve) irradiation of the reactive medium with the VUV surfatron-based source. It shows clear productions of C2 compounds, but also C3 and C4. This is confirmed by time-monitored spectra of the masses m/z=27 (figure 9) and 52 (figure 10) (possible molecules: HCN and C2N2 respectively) which reveal a sharp rise of the signals when the VUV source is turned ON (red lines).

The consumption of methane is moreover followed versus time (figure 8), showing a decrease of nearly 5% in 25 minutes.

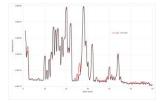
All the results are reproductible

#### **PERSPECTIVES**

The high flexibility of our VUV source will allow us to change the rare gas and thus the wavelength used to irradiate the reactive mixture. For example, helium provides 58-nm radiations and argon 104-nm ones.

Inside APSIS, the pressures can also vary as well as the CH<sub>4</sub> ratio.

Testing all those conditions will give us more information on the photochemical processes occuring in Titan's upper atmosphere and shed some light on prebiotic-like chemistry.





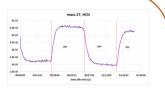
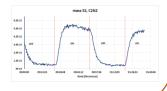


Figure 9 (top): time-monitored productions of HCN (m/z=27) showing reproductibility. The red lines show the moments the VUV source is turned ON.

Figure 10 (bottom): Same but for C<sub>2</sub>N<sub>2</sub> (m/z=52)



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