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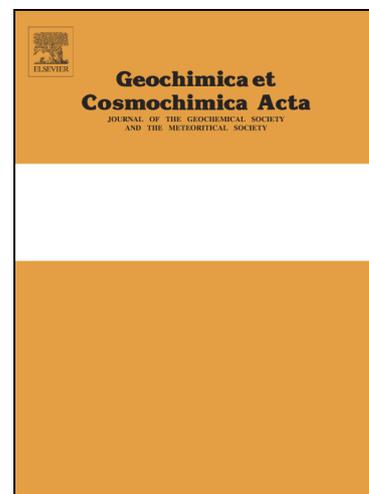
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## Capillarity-driven supersolubility in dual-porosity systems

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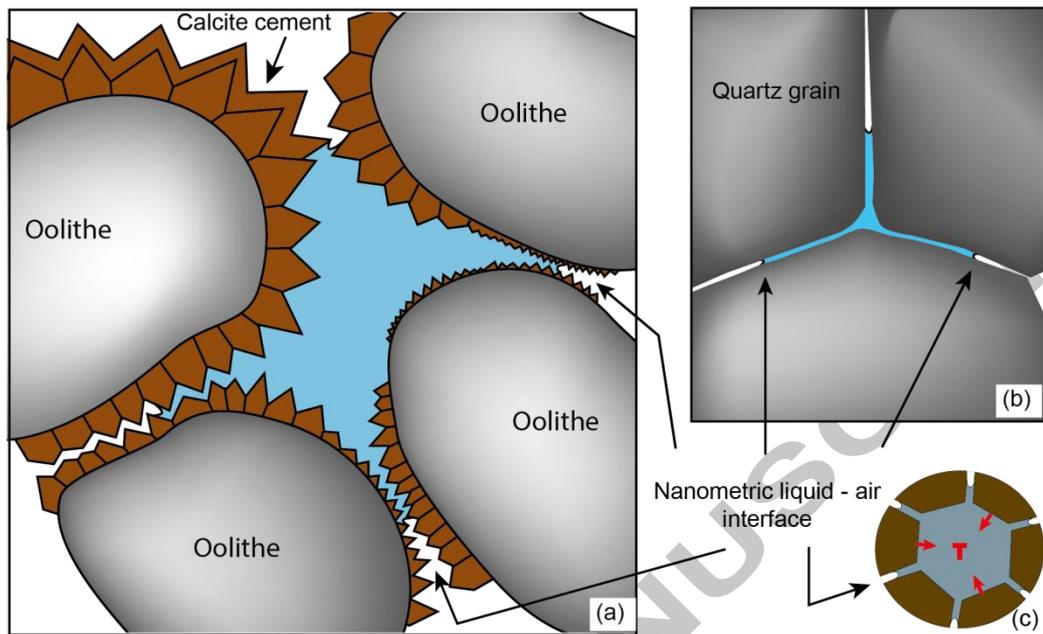
### Abstract

Capillary phenomena are widespread in unsaturated media such as soils, building stone, deep depleted aquifers and still gas storage/sequestration reservoirs. Processes linked to capillarity depend on pore size and environmental conditions such as relative humidity. When capillary forces are high, water's internal pressure can reach negative values representative of a tensile state. In those conditions, the capillary tensile water provokes compaction in moist granular materials. In geochemical terms, thermodynamics predict a significant effect on the solvent properties of tensile water. However, very little experimental work has validated thermodynamics models or shown the porosity arrangements that make capillarity a significant player to control the geochemical balance and poromechanics of porous media. In this study, we designed an experimental setup in glass microtubes ( $\varnothing$  200  $\mu\text{m}$ ) conducive to capillary tension. In the tubes, salts (halite and sodium sulfate) precipitated from an evaporating solution to build a strongly contrasted dual-porosity system. Large pore bodies ( $\varnothing$  200  $\mu\text{m}$ ) coexisted with nanometric pores that induced capillary tension in the whole volume. We observed mineral-liquid interactions (mass gain/loss, crystal shaping), especially how salt solubility changed as a function of capillary conditions. The water's tensile state increased salt solubility, changing the reactions' equilibrium constants at constant composition as predicted by thermodynamics. In addition, mineral grains showed clear evidence of poromechanical tensile stress, driving salts to move towards the interior of the tube or to crack. The results showed the potential significance of capillarity in heterogeneous porous media including nanopores. They should be considered in relation to how porous structures change, in the contexts of increasing droughts due to climate changes and increasing excavation work in deep sedimentary basins.

## 1. INTRODUCTION

Porous soils and rocks exposed to the atmosphere undergo weathering processes and erosion, which are heightened by frequent drying periods. Close to dry air, porous networks are characterized by the coexistence of air and water, and the fluid behavior is partly controlled by capillary forces. The so-called unsaturated zone (UZ) is widely present in near-surface environments, as soils and exposed rocks, as well as in engineering structures, monuments and artworks (Desarnaud et al., 2016). Global climate change, especially regarding humidity cycles (more frequent rainfall, flooding and extreme aridity, sea level rise), may directly impact behavior in the UZ and the associated weathering-erosion processes. Secondly, industrial processes such as gas injection, nuclear waste disposal and fluid removal drive deep geological formations to UZ-like states (e.g. André et al., 2014; 2015; Peysson et al., 2014; Ott et al., 2015), with possibly extremely dry (artificial) conditions.

Under drying conditions, water evaporates and the liquid–air interface retreats until it is retained in the thinnest pores by capillarity, where the internal liquid pressure decreases. The relative volume of capillary water results in poral geometry and dryness: for instance, multiscale porous networks retain larger volumes of capillary water, especially if large pores are connected through nanoscale pore throats (Fig. 1). In such a network, the nanometric capillary bridges in the pore throats generate high capillary forces that are transmitted to the whole volume of hydraulically connected water (outer pore throats and inner pore bodies) (Fig. 1). This situation can arise in sedimentary rocks (Fig. 1a), (e.g. Muller and McCauley, 1992; Wang et al., 2003; Anovitz et al., 2013) or in crustal rocks (Fig. 1b), (Holness, 1992; 1995).



**Figure 1. Dual-porosity systems. (a) in sedimentary oolitic rocks, (b) in crustal settings. (c) corresponding schematic view of a filled macropore surrounded by nanthroats (solid in brown, liquid in blue).**

Evaporation of in-pore water also leads to the precipitation of salt crystals initially dissolved in the poral solution. As they grow and accumulate in the poral network, those salt crystals can clog the pores and ultimately cause damages by exerting pressure against the pore wall (crystallization pressure, e.g. Correns, 1949; Scherer, 1999; Flatt, 2002; Benavente et al., 2004; Steiger, 2005a; b; Coussy, 2006; Shahidzadeh et al., 2012). Yet, in both arid environments and those with wetting and drying cycles, salts are recognized to be responsible for the alteration of porous media (e.g. Wellman and Wilson, 1965; Cooke, 1981; Goudie and Viles, 1997). While drying processes and evaporation have attracted a lot of attention, only a few studies have focused on the role of capillarity in weathering processes. However, capillarity is recognized to provoke the compaction of granular materials (e.g. Fraysse et al., 1999), possibly inducing cracks in clays (e.g. Weinberger, 1999) and cohesive sands (Chavdarian and Sumner, 2011). This effect can also be invoked to explain certain cracks in shotcrete (Ansell, 2010) and cementitious materials (e.g. Lagier et al., 2011) and deformation observed in soft porous material (Bouzzid et al., 2011b) and crystalline nano-channels (Tas et al., 2003). From the geochemical viewpoint, field observations suggest that a capillary state influences

pedogenetic evolution over millions of years, accumulating preferential anhydrous/hydrous mineral phases depending on climate cycles and mean pore sizes (e.g. Tardy and Nahon, 1985; Trolard and Tardy, 1989; 1988; Tardy, 1997). Capillary water's thermodynamic properties can be calculated either through a water activity framework (e.g., Trolard and Tardy, 1987; 1989; Wang 2003) or by a liquid pressure approach (Mercury and Tardy, 1997a, b; 2001; Mercury et al., 2003; 2004; Lassin et al., 2005). Both models predict how the geochemical mass balance will change in any UZ holding capillary water.

The effects of capillarity are size-dependent, and so may be significant only for nanometric pores. At this scale, the capillary water is under absolute negative pressure (or under capillary tension) and able to exert a tensile force (e.g. be responsible for granular stack compaction). Nanometric pores can be quantitatively important in sedimentary rocks. Macropores are connected through nanometric pore throats (Muller and Mc Cauley, 1992; Anovitz et al., 2013) as in Figure 1. Water retained in the large pore bodies surrounded by nanthroats can reach and sustain the same negative pressure as the nano-capillary water in the pore throats before water cavitation.

Our laboratory experiments studied the physical and chemical behavior of water-solid interactions in such situations. We developed a set-up based on an existing protocol (Fig. 2; Bouzid et al., 2011b; Røyne & Jamtveit, 2015) where a bimodal pore system conducive to capillary tension was built. Cylindrical microtubes were filled with a saline solution that evaporated in drying conditions. The solution became supersaturated with respect to its solid, which precipitated on both air-liquid interfaces (highest salt concentration). The salts grew until forming a cylinder that apparently clogged the tube, but a thin nanometric film persisted between the crystal and the glass wall (Fig. 2a). This film is assumed to have a zero contact angle with the silica walls, so the meniscus diameter is equal to the thickness of the open space. If the solution-air interfacial curvature became large enough to equilibrate with the imposed relative humidity (as expected from the Kelvin equation), capillary forces prevented evaporation and generated negative pressure in the entire trapped solution including the macrovolume. The entire volume of solution was in a tensile state (Fig. 2b), which is supported by the final bubble nucleation appearing inside the trapped macrovolume (Fig. 2c). The water under tension was in a metastable state (superheated) until the nucleation of vapor returned the system to its most

stable state (liquid–vapor assemblage at saturated vapor pressure). After bubble nucleation in the pore body, the liquid internal pressure is brought back at the atmospheric value. Consequently, there is no more pressure gradient across the liquid-air interface, which becomes flat (Fig. 2c).

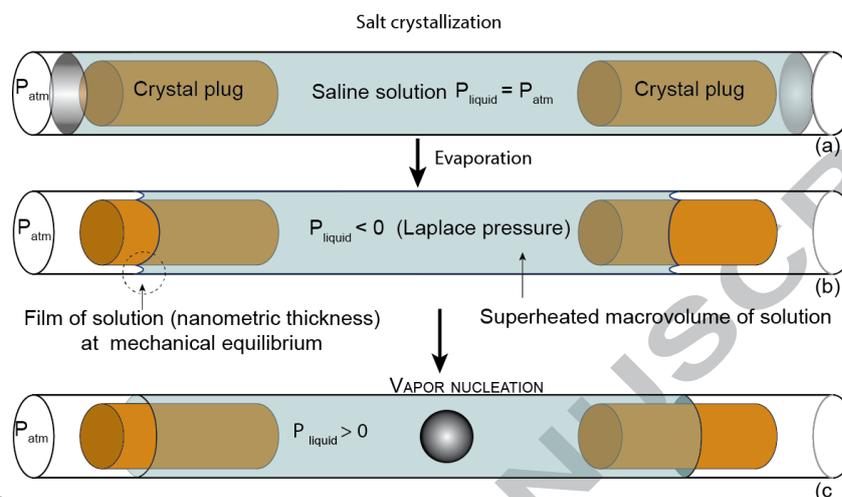


Figure 2. Main experimental steps (see text). (Modified from Bouzid et al., 2011b)

This method turns the initially monosized “saturated” microtube into an unsaturated bimodal pore system prone to high capillary conditions. Our goals were to experimentally validate the geochemical signatures of capillarity and to show that dual-porosity media can contain large capillary volumes.

## 2. MATERIAL AND METHODS

### 2.1 Materials

Micro-capillaries with 200  $\mu\text{m}$  internal diameter (VitroTubes<sup>TM</sup>) were filled with a saline solution ( $\text{NaCl}$  or  $\text{Na}_2\text{SO}_4$ ) that evaporated under controlled relative humidity.

$\text{NaCl}$  and  $\text{Na}_2\text{SO}_4$  aqueous solutions were prepared from NormaPur powder and distilled water. The  $\text{NaCl}$  solution was at  $5.5 \text{ mol.L}^{-1}$  (24.3 wt %) and had a relative supersaturation  $n/n_0 = 0.90$  (molality at saturation  $n_0 = 6.15 \text{ mol.L}^{-1}$  at  $20^\circ\text{C}$ ). The  $\text{Na}_2\text{SO}_4$  solution was at  $3.5 \text{ mol.L}^{-1}$  (33.3 wt %), so undersaturated with respect to thenardite (relative supersaturation  $n/n_0 = 0.64$  at  $20^\circ\text{C}$ ) and supersaturated with respect to mirabilite (relative supersaturation  $n/n_0 = 1.19$  at  $20^\circ\text{C}$ ). The saline solutions were filtered with  $0.450 \mu\text{m}$  filters (VWR® syringe filters) to avoid impurities and micro-

crystallites. This means that there are no or very few potential seeds in the solution, so that during evaporation, the solution will reach a high supersaturation before crystals nucleate (Desarnaud et al., 2014).

The evaporative demand, leading to evaporation or capillary conditions, was set by controlled relative humidity (RH) in a climate chamber (HPP108, Memmert technology). The preferred RH was 48%, to be compared to the equilibrium RH of a NaCl saturated solution (75.3%, Greenspan, 1977) or of a Na<sub>2</sub>SO<sub>4</sub> solution (93%, Winston and Bates, 1960). This RH gradient enables a moderate drying rate promoting the precipitation of massive salt (Desarnaud et al., 2014) when the appropriate supersaturation was reached, at constant temperature. The microcapillaries were regularly observed using an optical microscope (Leica DM2500) in a room thermostated at 20°C. Observations were made in transmitted light using lens with 5x, 10x, 20x and 50x magnification. Pictures were captured with a Leica DFC295 digital camera allowing image resolution of 3 megapixels (2048 x 1536) and pixel size of 1.2 µm x 1.2 µm with a 5x magnification lens. To avoid any heat transfer, we reduced exposure to microscope lights to the minimum possible, so we can consider heating due to light as negligible.

## 2.2 Methods

Three types of experiments (Fig. 3; also supp. info.) were carried out, differentiated according to:

- two lengths of liquid column trapped in between the crystal plugs: long and short, corresponding to large and small volumes of tensile water;
- two types of initial solutions, based on NaCl and Na<sub>2</sub>SO<sub>4</sub> electrolytes, because the solvent properties of the associated crystals phases are different when the water is in a tensile state.

The NaCl-short solution systems (Fig. 3a-d) were made from a highly supersaturated solution, using plasma-cleaned tubes (15 min) and filtered solution. This made the halite nucleate inside the solution and grow rapidly on both sides with a symmetric hopper (skeletal) shape (Sunagawa, 1999; Desarnaud et al., 2014) (Fig. 3a). First, when the supersaturation was still high, several cubes of halite nucleated and crystallized on each other, with growth parallel to the tube. As the supersaturation decreased, the salt continued to grow at each side preferentially onto the inner wall of the tube. As the

solution kept on evaporating, the salt continued to grow on each side and not in the middle, because at this point the solution was trapped and did not evaporate. This explaining the “bow-tie” shape (Fig. 3b-c). The solution on each side continued to evaporate until reaching first one side of the crystal (Fig. 3c) then the other (Fig. 3d). The liquid air interface stopped receding since no air entered in the trapped solution. We built 26 “bow-tie” systems.

For the long-NaCl solution systems (Fig. 3e), the NaCl solution was left to evaporate until reaching high supersaturation, without any crystal nucleation or growth. Then we broke the tube on each side close to the air–water menisci. This induced rapid evaporation at both liquid–air interfaces and led to the precipitation of two opposite crystals along the interfaces. The two crystals grew rapidly and apparently clogged the tube, trapping the large volume of solution (from 6.9 to 51 mm long, so from 0.157 to 1.88 mm<sup>3</sup> in volume). As in the bow-tie shape systems, the trapped solution did not evaporate. Twelve of these systems were built.

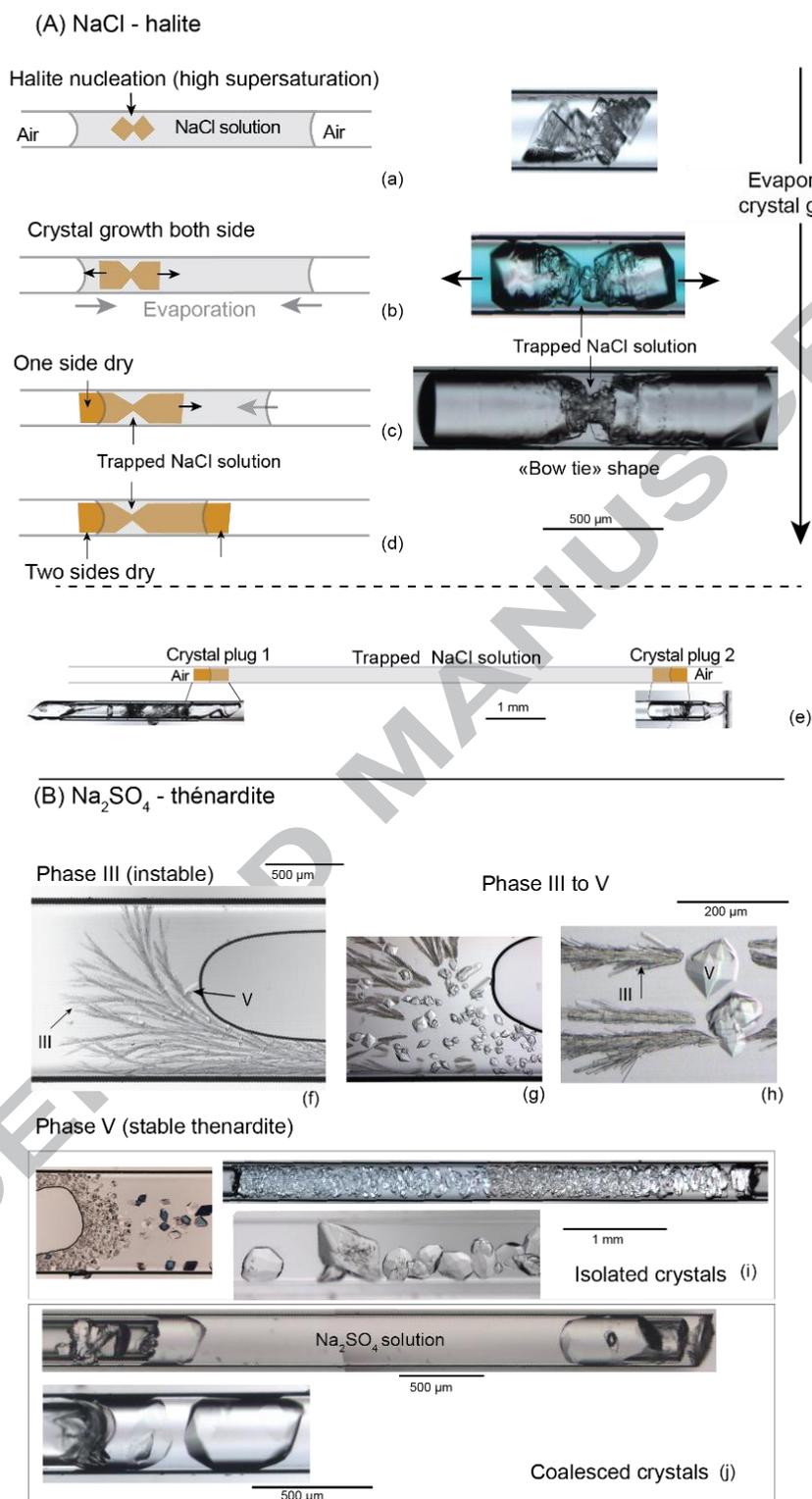


Figure 3. Crystallization sequences to build double plug systems (see text).

Concerning Na<sub>2</sub>SO<sub>4</sub>, no useful distinction was made by tube length, so the experiments were performed irrespective of the trapped water volume. The protocol consisted of filling the tube with a

33.3% wt solution, undersaturated with respect to thenardite (anhydrous stable phase) and supersaturated with respect to mirabilite (decahydrate stable phase). After evaporation, the salt first crystallized into a dendritic metastable phase III (Fig. 3f) then transformed into stable phase V, thenardite (Fig. 3g-h), as commonly observed (Steiger and Asmussen, 2008; Shahidzadeh and Desarnaud, 2012). The crystallization sequence resulted in numerous thenardite crystals in solution (Fig. 3i), which coalesced to form plugs (Fig. 3j). The formation of massive plugs was encouraged by using highly saline solution (33.3% wt) introducing high salt mass. Ten such systems were built.

### 3. RESULTS

The dual-porosity systems were built in 48 tubes (table 1) either with one bow-tie shaped halite crystal (Fig. 3d), or with two distinct crystals (halite or thenardite). In each system, after the liquid–air interface receded and reached the external side of a crystal (Fig. 2b), we observed no further air entry in the trapped macrovolume. The liquid–air interface remained in the narrow space between the salt plugs and the glass wall and stopped evaporating. After one external side of one crystal was dry (and only after), we observed different reactions in the trapped volume: salt dissolution, gas nucleation and solid displacement. Over time (up to 2 years), we carefully monitored the events occurring within the trapped solution and/or affecting the crystal plugs.

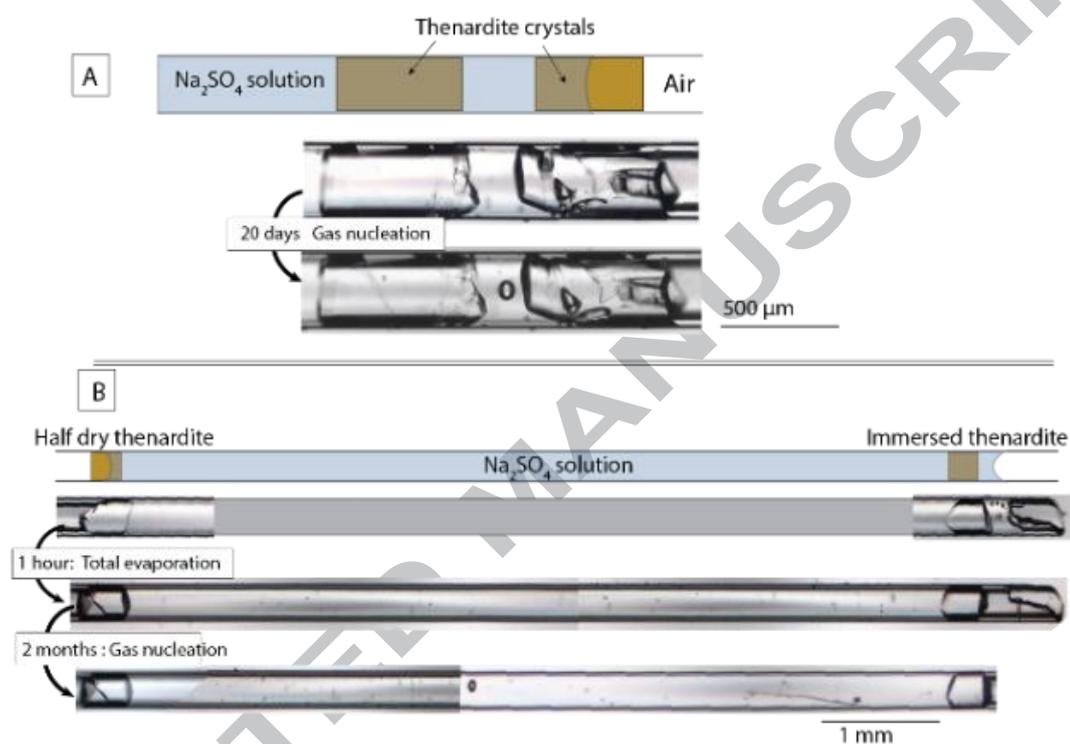
Exp type	Number of exp	Solid motion	Salt dissolution	Gas nucleation	Time before gas nucleation	No final gas nucleation	
						No changes	Only solid motion
NaCl-short	26	0	16	20	2 days to 3 months	6	
NaCl-long	12	4	1	5	4 days to 3 months	5	2
Na <sub>2</sub> SO <sub>4</sub>	10	0	0	10	20 days to 3 months	0	
Total	48	4	17	35	-	11	2

**Table 1. Summary of dual-porosity systems and major observables (also suppl. info.).**

#### 3.1 Gas nucleation

The most common reaction, observed in 35 systems, was the formation of a gas bubble in the trapped solution after a certain amount of time – from several hours to months – after at least one of

the external sides of the salt(s) was dry. This gas nucleation has already been observed and was not the result of air entry (Bouزيد et al., 2011b). Figure 4 illustrates two experiments in which gas nucleated in a solution trapped between two thenardite crystals ( $\text{Na}_2\text{SO}_4$ ). In the first experiment, gas nucleated 20 days after the external side of the right crystal dried (Fig. 4A) and in the second experiment, gas nucleated two months after the external sides of both crystals dried (Fig. 4B).

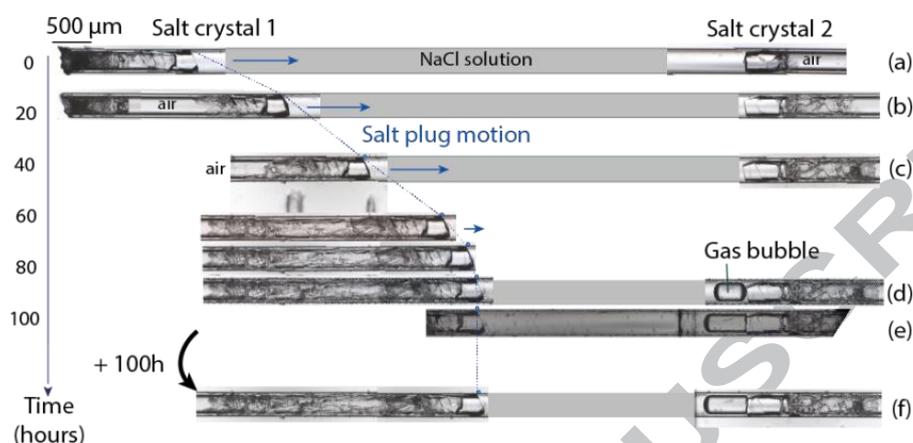


**Figure 4. Gas nucleation in a  $\text{Na}_2\text{SO}_4$  solution trapped by two thenardite crystals. (A) 20 days after one external side of the salt was dry (B) 2 months after both external sides of the salt became dry.**

We observed gas nucleation in the trapped solution in bimodal pore systems built with  $\text{NaCl}$  (halite) and  $\text{Na}_2\text{SO}_4$  (thenardite). In each of the 10 systems built with thenardite, gas nucleated from 20 days to 3 months after at least one external side of the salts dried. In systems built with halite crystals, we observed gas nucleation in 25 of the 38 experiments: in 20 of the 26 bow-tie shape systems (Fig. 3a-d), and in 5 of the 12 systems built with two distinct salts (Fig. 3e; see also supplementary information).

### 3.2 Inward movements of salt crystals

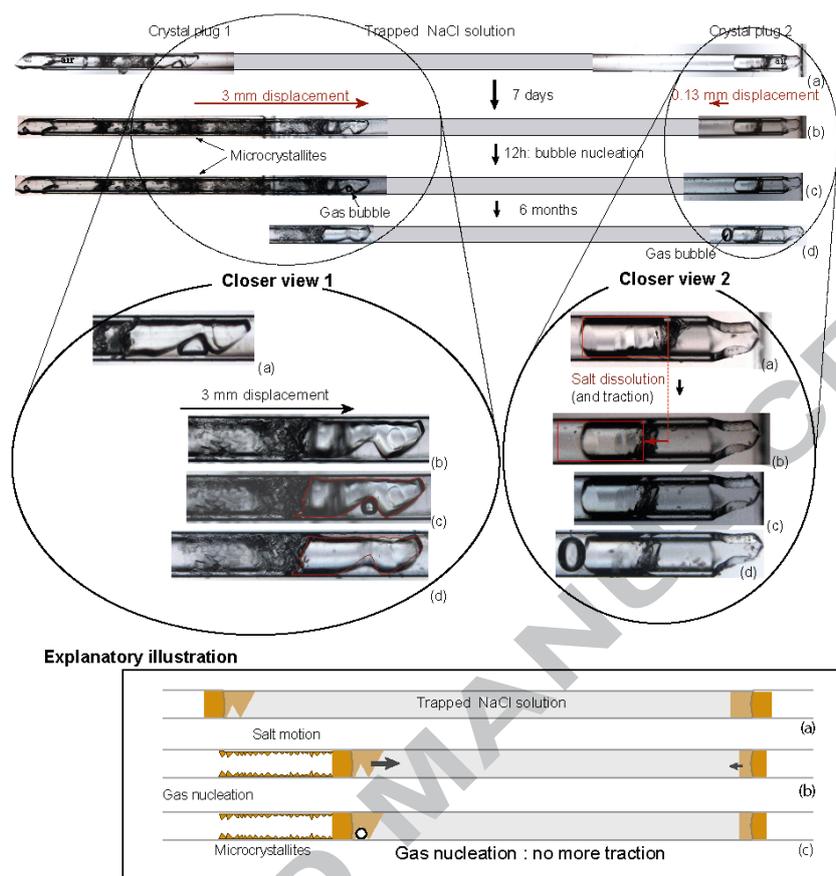
With two of the NaCl-long systems, crystal plugs were seen to move (Fig. 5).



**Figure 5. Salt plug inward motion, evidenced with the large displacement of the left salt crystal that comes to an end with bubble nucleation.**

In both cases, we observed the inward motion of one or both salt plugs until a gas bubble appeared in the trapped solution. In the experiment shown in Fig. 5, the crystal on the left slid over 3.3 mm in four days (Fig. 5b-c). When a gas bubble appeared on the right (Fig. 5d), the crystal stopped and never moved again. In the experiment shown in Fig. 6, both salts moved inward for seven days (Fig. 6, step a to c) before a gas bubble nucleated (Fig. 6, step d). The salt on the left moved over 3.3 mm while the opposite one (on the right) moved over 0.13 mm (see the closer view 2, Fig. 6).

No plug motion was observed in any of the 10 Na<sub>2</sub>SO<sub>4</sub> based systems.



**Figure 6.** Salt traction during the 7 days after both salt plugs precipitated, trapping the saline solution. **At the top:** pictures of four stages of the experiment. **Below:** closer views of the salt plugs: the crystal on the right (closer view 2) first dissolved (step b) and then slightly reprecipitated (step d). **At the bottom:** illustration of the first three stages of the experiment for a better understanding. **Step (a):** salt plugs precipitated and entrapped the saline (NaCl) solution. **Step (b):** after 7 days, the crystal on the left has moved inward over 3.3 mm, while the one on the right has moved over 0.13 mm and been dissolved (see closer view 2). **Step (c):** gas nucleated in the trapped solution and the salts stopped moving and dissolving.

### 3.3 Dissolution of halite

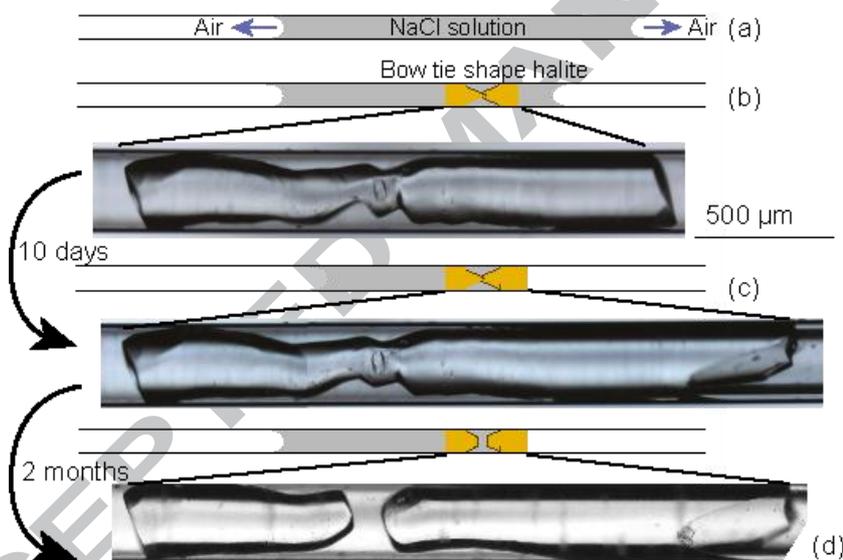
#### 3.3.1. Dissolution and traction

The moving crystal on the right in the NaCl-long experiment had been partially dissolved before the gas appeared (Fig. 6A, closer view 2). The simple cylindrical geometry of the system allowed us to measure the volume of the trapped solution and the total mass loss. The dissolved length of halite was  $130 \pm 13 \mu\text{m}$ , corresponding to a volume of  $4.1 \cdot 10^{-3} \pm 0.4 \cdot 10^{-3} \text{ mm}^3$ , and therefore a mass loss of  $8.9 \pm 0.9 \mu\text{g}$  (calculated with a halite density of  $2.17 \text{ g/cm}^3$ ). The distance between the two salt plugs before (Fig. 6A, step a) and after crystals moved (Fig. 6A, step b) was respectively  $10 \pm 0.2 \text{ mm}$  and  $6.6 \pm$

0.2 mm, corresponding to a solution volume of  $0.32 \pm 0.01 \text{ mm}^3$  and  $0.21 \pm 0.01 \text{ mm}^3$ . Furthermore, a close examination of the same halite plug after the bubble nucleated showed clear lengthening over time, evidencing a reprecipitation process: the process appeared reversible.

### 3.3.2. Bow-tie shaped systems

Whereas salt dissolution was only observed in one experiment built with two distinct salts (large double plug systems, Fig. 3e, j), significant salt dissolution was observed in 16 of the 26 experiments built with bow-tie shape halite (Fig. 3A). Dissolution affected the salt in contact with the trapped solution at the thinnest part. In most cases, the bow-tie crystal was separated into two pieces and both pieces became rounded (Figs. 7 and 8, steps c and d).

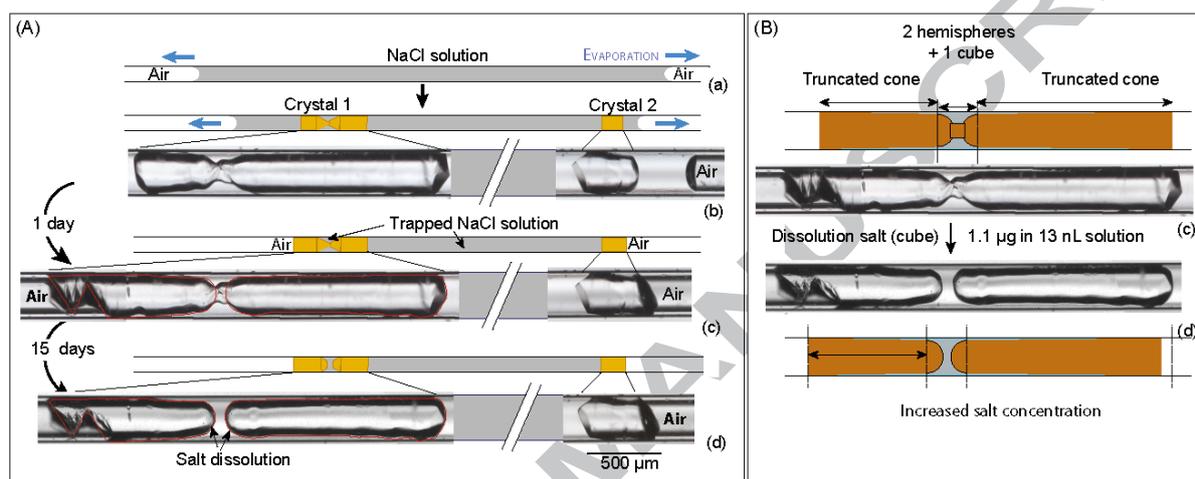


**Figure 7. (A) Salt dissolution in the trapped solution after one external side of the crystal (on the right) encounters air – photos and interpretative sketch. (a) The NaCl solution evaporates. (b) Bow-tie shape crystal growing in the evaporating solution. (c) After 10 days, the right side of the crystal is dry. (d) Dissolution of the thinnest part of the salt in contact with the trapped solution.**

Twenty of the 26 experiments built with bow-tie shaped halite showed gas nucleation after a certain period, between 5 days and several months (section 3.1). Gas nucleated in each of the 16 cases displaying salt dissolution, and there were no cases where salt dissolved after gas nucleation.

Dissolution in bow-tie shaped crystals lasted between 5 and 45 days. In most of the cases dissolution time was 10 to 15 days.

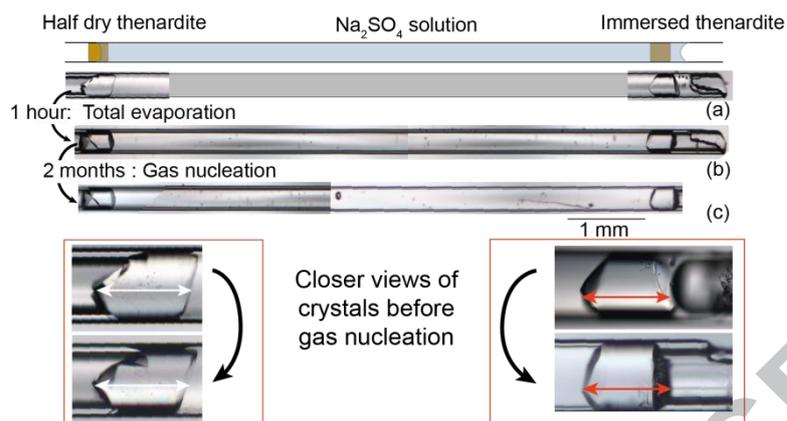
Due to the complex geometry of the bow-tie halite crystals, we could only calculate the volume of solution and the mass loss in one case (Fig. 8B). The volume of the trapped solution was estimated at  $13.10^{-3} \pm 1.3.10^{-3} \text{ mm}^3$ , and the mass loss of solid salt at  $1.1 \pm 0.1 \mu\text{g}$  (Fig. 8B).



**Figure 8.** (A) System with two trapped compartments of saline solution. (B) Estimation of salt loss and volume of solution outlined by the red line drawn around the crystal (c and d).

### 3.3.3 The case of thenardite

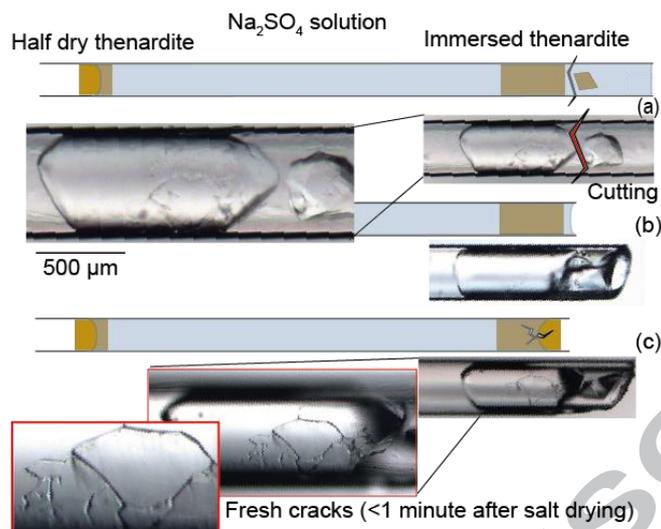
In  $\text{Na}_2\text{SO}_4$  systems, we observed neither dissolution nor movement of the thenardite crystal. However, as in NaCl systems, a gas bubble appeared in the trapped solution, from one day to months after one or both salt plugs dried (Fig. 9). A closer view of both surrounding salt plugs did not show any mass change with respect to the initial situation: no significant dissolution occurred during those two months.



**Figure 9.** Absence of visible salt mass change in  $\text{Na}_2\text{SO}_4$  double plug systems. Double arrows (red and white) on the closer view show that salt mass had not changed (one color = one size).

### 3.4 Plug cracking

Another interesting observation was made with  $\text{Na}_2\text{SO}_4$  double plug systems, at the critical moment when one salt plug dries. In this experiment we cut the tube (Fig. 10, step a), enhancing the evaporation of the saline solution. After around ten minutes, the liquid–air interface reached the salt plug (Fig. 10, step b) and became confined in the small space between the salt plug and the tube wall. As the interface continued to recede between the salt and the tube wall, cracks appeared and propagated on the crystal surface along preexisting discontinuities (probably the mark of crystal coalescence).



**Figure 10.** Crack development observed after a crystal plug dried (a) The tube is cut close to an immersed salt plug. (b) The liquid-air interface retreats and reaches the crystal which becomes dry (c) Cracks grow from the discontinuities on the crystal surface (see closer view in (b)).

## 4. DISCUSSION

### 4.1 Tensile state and tensile stress

#### 4.1.1 Gas nucleation and capillary film

The presence of a capillary film along a nano-scale space between the solid plugs and the tube walls is a first assumption of this work. We aspire to measure the properties and/or thickness of this film but have not achieved this yet. The Young-Laplace law (suppl. info.) gives the pressure drop across a curved interface between two phases. With concave air-liquid interfaces, the liquid pressure is lower than the air pressure, and increases as interface size decreases. This interface is itself related to the pore size by the value of its contact angle with both supporting solids (suppl. info.). As a second assumption in the calculations, the contact angle between the liquid and the solid host is assumed to be zero.

The thermodynamic driving force towards capillarity comes from air's RH when below the saturated vapor pressure (suppl. info.). If the water-air surface is flat, low RH drives evaporation that increases the local air vapor pressure. In small pores, low RH can drive the liquid to decrease its

internal pressure (the water is in capillary state). If the local pore size is sufficiently small, the liquid–air interfacial curvature becomes large enough to equilibrate water and dry air (Kelvin equilibrium is reached). The liquid pressure can decrease according to the Young-Laplace law (eq (2), suppl. info.).

Consequently, the decrease in liquid pressure, driven by dry air, is controlled and limited by the size of the nano-throats and is identical in the entire hydraulically connected volume. Yet the macropores are so large that they cannot prevent the boiling of tensile water. The formation of a sphere of vapor of radius  $r$  brings the system to its most stable state which gains energy equal to  $\frac{4}{3}\pi r^3(P_{sat} - P)$  with  $P$  the pressure of the tensile liquid and  $P_{sat}$  the saturation vapor pressure). However, the creation of the vapor-liquid interface of radius  $r$  requires an activation energy equal to  $4\pi r^2\gamma$  (with  $\gamma$  the liquid–vapor surface tension). According to the Classical Nucleation Theory (CNT), the competition between these two opposite effects results in an energy barrier  $E_b$ :

$$E_b = 4\pi r^2\gamma - \frac{4}{3}\pi r^3(P_{sat} - P) = 16\frac{\pi}{3}\frac{\gamma^3}{(P_{sat} - P)^2} \quad (1)$$

which is the minimum energy required to create a sphere of vapor of critical radius  $r$  (e.g. Debenedetti, 1996, Caupin and Herbert, 2006). In nanothroat, the size is too small for the vapor nucleus to reach the critical radius required for vapor nucleation. In contrast, the adjacent connected macropore is much larger than the critical radius, enabling vapor nucleation at any time. Before that, the occluded tensile water is metastable (superheated) during a limited lifetime that decreases with the superheating intensity, which is directly depending on  $P_{sat} - P$ .

During the course of our experiments, we observed vapor nucleation in 35 on 48 microtubes after a superheating period ranging for a few days to years.

At the nanoscale, repulsive forces are caused when a fluid sandwiched between two solids undergoes the overlapping of surface force fields emanating from the facing solids (e.g. Derjaguin et al., 1987). This disjoining effect accounts for the open film, even in the presence of a supersaturated solution (e.g. Desarnaud et al., 2016). However, impurities, wall roughness and excess supersaturation, for instance, could initiate precipitation of microcrystals that would seal the open film and totally clog

the tube. This process could explain the absence of any change in the system, including gas nucleation, in six NaCl-short and 5 NaCl-long experiments (Table 1; supp. info). Water in the macrovolume would simply be trapped, without any capillary forces to put it under tensile stress, and therefore without gas nucleation, enhanced dissolution, or plug motion. It is worth saying that two NaCl-long experiments first showed slow inward halite plug motion for 3.5 and 7 months. Once the plugs stopped, no gas nucleation happened in the following year, meaning that the trapped macrovolume of solution can maintain a tensile state for a long time.

#### 4.1.2 Mechanical effects

The inward motion of salt plugs (Figs 5 and 6) demonstrates that the crystals did not obstruct the tube and were mechanically free to move. This is consistent with the macrovolume of solution having a tensile state before vapor nucleation. As vapor nucleated, the stretched water relaxed its tension, becoming unable to exert any tensile forces on the salts. In the two experiments where salt plugs moved inward, the two salt plugs moved differently (Fig. 5) or only one salt plug moved (Fig. 6). The liquid must exert an isotropic and homogenous tension in the macrovolume, but the friction forces between the glass wall and each solid plug should have been different, probably because each crystal had its own surface roughness.

The plug cracks observed at the thenardite surface (Fig. 9) are assumed to be the result of capillary tension exerted by the nanometric liquid–air interface. Cracks appeared as the nanometric interface receded between the salt and the tube wall, several minutes after we cut the tube. Therefore, the cracks were not related to the mechanical procedure of cutting the tubes. These cracks revealed the mechanical constraints exerted by the liquid against the solid surface. Yet the tensile stress in liquid is isotropic, so the mechanical effects depend primarily on the strength of the solids themselves. The glass wall was probably unaffected because of its high mechanical resistance. The thenardite crystals may have cracked because of the preexisting defects linked to the coalescence formation process (Fig. 9). Considering texture heterogeneity of pore walls in natural rocks, especially discontinuity in the structure or dislocations, even low tension can have strong mechanical effects.

We used two easy methods to ensure that the  $\text{Na}_2\text{SO}_4$  that crystals were thenardite (anhydrous crystals) and not mirabilite (decahydrate crystals). First, mirabilite rapidly dehydrates in microcrystals of thenardite in contact with dry air with RH of 93% or less (equilibrium vapor pressure of mirabilite). This event is very characteristic and easily recognizable. As the experiments were conducted at 48% RH, mirabilite must dehydrate, so we can directly identify the crystals during the experiment. Secondly, mirabilite dissolves at  $32.4^\circ\text{C}$  (e.g. Steiger and Asmussen, 2008). To be sure that the salt plugs were thenardite, at the end of experiments we heated the sample in a Linkam stage. As the crystal plugs did not dissolve above  $32.4^\circ\text{C}$ , we confirmed that the crystals observed were indeed thenardite. This has been confirmed using XRD analysis on two tubes containing  $\text{Na}_2\text{SO}_4$  salts that we identified as thenardite from morphology, behavior in dry environment, and heating.

## 4.2 Supersolubility effect

The dissolution of halite demonstrated that the trapped solution was becoming undersaturated with respect to halite. As temperature and composition remained constant, the change necessarily relates to the reaction's equilibrium constant. This change is attributed to the liquid pressure decreasing toward negative values, which is consistent with the other observations, like gas nucleation and salt plug traction. The tension was imposed by capillary forces settled by one or two nanometer scale menisci surrounding the macrovolume of solution. Thermodynamics involving capillarity (Mercury and Tardy, 1997 a, b; 2001; Mercury et al., 2003; 2004; Lassin et al., 2005; see also supp. info) predict how the equilibrium constant of any reaction changes when the liquid pressure changes.

$$\ln \frac{K_P}{K_0} = - \frac{\int_{P_0}^{P_{cap}} \sum V^{\circ}_{dissolved\ ions} dP + n \cdot \int_{P_0}^{P_{cap}} V^{\circ}_{H_2O} dP}{RT} \quad (2)$$

$K_P$ ,  $K_0$  are the equilibrium constant for reactions interacting with capillary and bulk liquid respectively,  $V^{\circ}$  is the ionic volume of the subscribed phase at infinite dilution,  $n$  is the number of  $\text{H}_2\text{O}$  molecules involved in the reaction (from 0 for halite to 10 for mirabilite dissolution),  $R$  is the gas constant and  $T$  is the temperature in kelvin. The equilibrium constant  $K_P$  in capillary conditions was calculated

considering that the chemical potentials of solids kept their values at atmospheric conditions (suppl. info.).

The increase of solubility of halite, thenardite, and mirabilite can be calculated according to equations (2), (3) and (4). For halite, the solubility  $S$  is the square root of the solubility product  $K$ , so

$$\ln \frac{S_P}{S_0} = 0.5 \ln \frac{K_P}{K_0} \quad (3)$$

with  $S_P$  the halite solubility in capillary conditions (negative pressure  $P$ ) and  $S_0$  the halite solubility at atmospheric pressure. For thenardite and mirabilite, the solubility  $S$  is the cubic root of the solubility product  $K$ , so

$$\ln \frac{S_P}{S_0} = \frac{1}{3} \ln \frac{K_P}{K_0} \quad (4)$$

The solubility ratio  $\frac{S_P}{S_0}$  of halite, thenardite and mirabilite are reported in Fig. 11 as a function of capillary negative pressure.

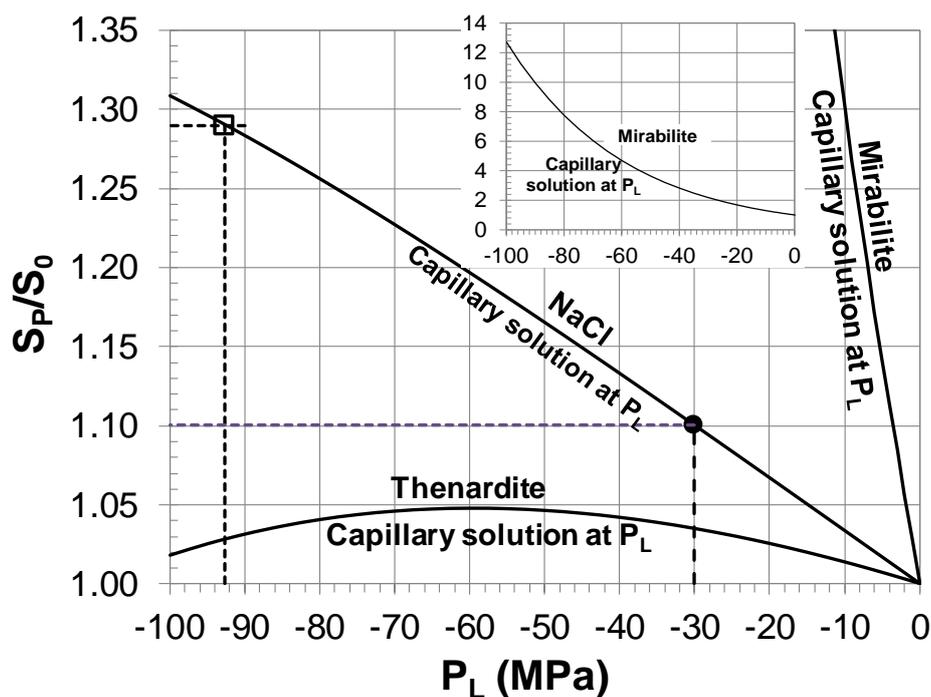


Figure 11. Thermodynamic change in equilibrium constant with increasing tension in the interacting aqueous solution. The black circle and hollow square show the positions of the extreme liquid tensions accounting for the enhanced dissolution respectively observed with NaCl-long (Fig. 6, 30 MPa) and NaCl-short (Fig. 8, 93 MPa) experiments.

The geochemical behavior can be readily interpreted considering the volume properties of the solutes under tension (Mercury et al., 2003; also suppl. Info.).

By calculating the mass of dissolved halite, it was possible to calculate the tension required for the new equilibrium constant. For that, we assumed that the halite + liquid system had reached the new equilibrium imposed by capillary tension. If so, the increased concentration we calculated with the mass loss corresponds to the increase in halite solubility. In some experiments (Figs 7 and 8), the dissolution stopped largely before gas nucleation, supporting this assumption. In others, gas nucleated while the solid plugs were dissolving, meaning our calculations gave a minimal tension value, probably underestimated.

In two experiments (Figs 6 and 8), the tension could be estimated because of the simple geometry of crystal plugs and liquid macrovolume. In the NaCl-long experiment (Fig. 6),  $8.9 \pm 0.9 \mu\text{g}$  of halite dissolved in a volume of solution of  $0.21 \pm 0.01 \text{ mm}^3$ . Considering the assumptions stated above, this

corresponds to an increased mass concentration of  $43 \pm 7 \mu\text{g}\cdot\text{mm}^{-3}$  which corresponds to an increased solubility of  $12 \pm 2\%$  ( $\frac{S_P}{S_0} = 1.12 \pm 0.02$ ). The first assumption was that activity coefficients do not change with the increasing concentration linked to the higher solubility. The second assumption is that the system reached equilibrium during the observation time: both the macropore volume and the halite plugs retained their characteristic sizes. The thermodynamic model predicted a tension of  $36 \pm 6$  MPa (Fig. 11, black circle shows the minimal tension 30 MPa). To reach that tension, the liquid–air interface must have a curvature radius of  $2.4 \pm 0.4$  nm and a thickness of  $4.8 \pm 0.8$  nm if the contact angle is equal to zero.

In one NaCl-short experiment (Fig. 8),  $1.1 \pm 0.1 \mu\text{g}$  dissolved in  $13 \times 10^{-3} \pm 1.3 \times 10^{-3} \text{mm}^3$  of solution, corresponding to an increased mass concentration of  $86 \pm 16 \mu\text{g}\cdot\text{mm}^{-3}$ . This corresponds to an increased solubility of  $24 \pm 5\%$  ( $\frac{S_P}{S_0} = 1.24 \pm 0.05$ ) due to a tension of  $75 \pm 18$  MPa (Fig. 11, hollow square shows the maximum tension 93 MPa). To reach that tension, the liquid–air interface must have a curvature radius of  $1.4 \pm 0.3$  nm and a thickness of  $2.8 \pm 0.6$  nm if the contact angle is equal to zero. In this case, the solid dissolution stopped two weeks before gas nucleated, in close agreement with the assumption that the system had reached a new equilibrium state.

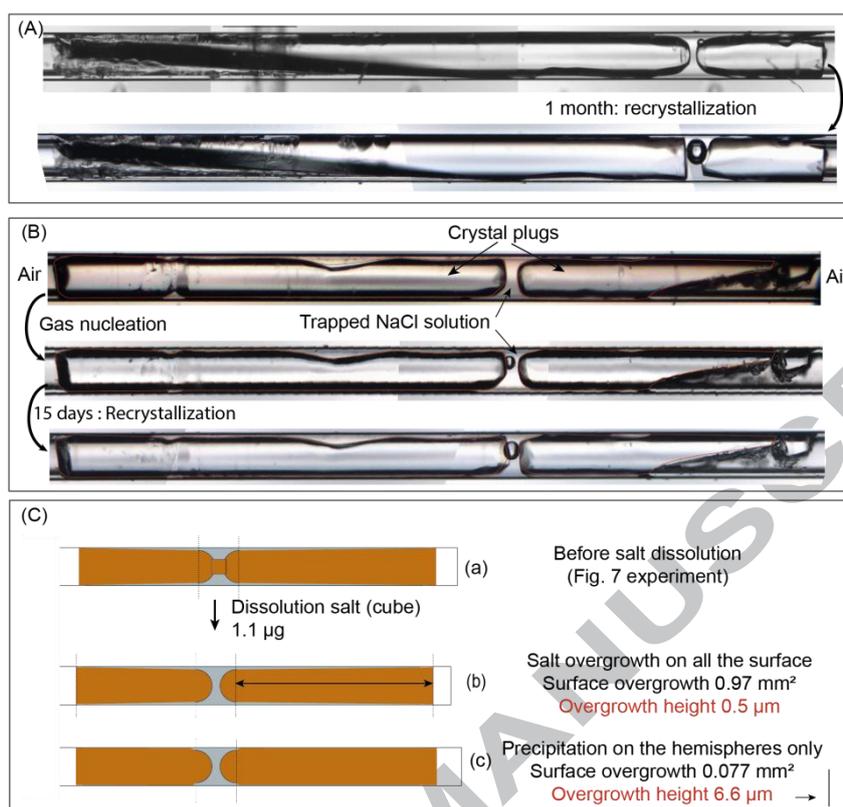
In most of the NaCl-short experiments, the amount of dissolved salt could not be calculated because of the complexity of the crystal shape and of the liquid volume. However, the increased solubility was clearly evidenced by the dissolution of the central part of the mono-solid, leading to two distinct halite plugs. This observed dissolution of the central part of the mono-solid is also consistent with minimizing surface energy of the solid-solution interface.

It is noteworthy that water and aqueous solutions can sustain even higher tensions (100 MPa and beyond) in closed microcavities than the presently recorded capillary tensions (30 to 93 MPa) (e.g. Roedder, 1967; Green et al., 1990; Zheng et al., 1991; Alvarenga et al., 1993; Shmulovich et al., 2009; Pallares et al., 2016; Qiu et al., 2016). Interestingly, Chen et al. (2016) developed an ink-bottle pore geometry, by bonding a nanoporous silicon membrane onto a glass substrate patterned with large voids. They measured a maximum tension sustainable by the stretched superheated water in their macrovolume from 20 to 30 MPa, matching the present values.

Regarding thenardite ( $\text{Na}_2\text{SO}_4$ ), we never detected salt dissolution, while the liquid's tensile state was illustrated by the final water cavitation. The absence of dissolution can be interpreted by the limited increase of thenardite solubility with tension, which reaches a maximum of 5% at -60 MPa before decreasing (Fig. 11).

### 4.3 Re-equilibration after vapor nucleation

After vapor nucleation, the liquid returned to its most stable state (e.g. Caupin and Herbert, 2006) and the solid salt solubility to its initial value. Consequently, after bubble nucleation, the aqueous solution is instantaneously supersaturated with respect to the solid, and recrystallization is expected to restore the new (though prior) equilibrium. However, especially in the bow-tie shape experiments, the crystal could not recover its initial shape because the secondary crystallization occurred through epitaxial overgrowth on the existing mineral. Considering the salt quantity in excess and the crystal surface area in contact with the solution, the thickness of the overgrown layer would be around  $1 \pm 0.5 \mu\text{m}$  along the whole surface area of the solid, or  $7 \mu\text{m}$  with growth restricted to the two curved hemisphere surfaces (Fig. 12C). In both cases, we were not able to visualize the reprecipitated mass under the microscope (Fig. 12, see black arrow: a line  $7 \mu\text{m}$  thick was drawn to appreciate the order of magnitude). In conclusion, salt dissolution was visible and calculable because it only occurred at the thinnest part of the salt where crystal morphology includes more kinks and steps whose surface energy is higher than a flat surface (Chernov, 2010). On the contrary, when the salt reprecipitated to reach the new equilibrium, it was deposited as a thin layer along the whole surface of the existing salt in contact with the saline solution. Consequently, the thickness of this secondary layer was so thin that we could not see it with an optical microscope. However, the reprecipitation materializes in the reshaping: the halite plugs changing from a round shape to halite's typical cubic habit (Fig. 12A, B).



**Figure 12.** Changes in of two salts after the bubble nucleates, relaxing the solution's superheating state. The vertical black line at the bottom right is 7 μm thick and is drawn to show the maximum expectable thickness of the reprecipitation layer (see text).

## 5. CONCLUSION

This experimental approach illustrates that capillary tension up to tens of MPa can arise in a macrovolume of trapped aqueous solution. The capillarity tension is associated with a bimodal porosity synthesized in an initially monosized tube where sodium sulfate and halite precipitated in an evaporating solution. The dry air conditions additionally allow a nanometric liquid–air curved interface to exist in the thinnest part of the porous domain. These dual-porosity systems enabled us to monitor clear and significant mechanical and geochemical re-equilibration when put under capillary conditions. Increasing solubility, inward motion of grain crystals and crystal cracking were carefully observed in various situations and displayed convergent behaviors. Furthermore, a supersolubility effect can be readily interpreted using a thermodynamic framework that resulted in calculated tensions ranging from 30 to 93 MPa.

The thermodynamic framework was previously used to predict how capillarity might impact the geochemical budget along UZ either through indirect clues (e.g. Tardy and Nahon, 1985; Trolard and Tardy, 1987; 1989; Tardy, 1997) or using thermodynamic predictions (e.g. Mercury and Tardy, 1997; Mercury et al., 2003; 2004; Lassin et al., 2005; Pettenati et al., 2008; Pedretti et al., 2015). Recently Lassin et al. (2016) demonstrated experimentally on CO<sub>2</sub>-water equilibria that these predictions matched measurements. This work establishes that mineral-liquid interactions are significantly impacted and are associated with various mechanical effects that lead to a great variety of transport/reactivity couplings in geological systems.

Interestingly, the experiments we proposed showed both geochemical and poromechanical effects of the capillarity driving force, associated with the existence of a secondary thin porosity in a complex multiphasic system. The tensile stress can modify the pore architecture by displacing in-pore mineral grains, but also may play a role in terms of permeability with crack growth. However, the signatures depend on the chemical systems involved, and that substantially widens the potential variety of patterns in a multi-mineral porous media.

Conversely, our experiments demonstrated that salts precipitating within macro-scale pores or channels can build the required topology for the capillary tension to exist. This means that porous media cementation and the existence of nano-scale throats must be considered for transport property variation over time but should also be carefully examined to enable conclusions on the water-rock interaction processes. The required nano-scale pores are indeed largely encountered in geologic media, especially in sedimentary rocks where large pore bodies are connected through nanometric pore throats (e.g. Muller and McCauley, 1992; Wang et al., 2003; Anovitz et al., 2013). The ratio of nanopores (let's say, 1 to 100 nm) to the total porosity can be quite large up to, for instance, 10-40% (Görres et al., 2000), 20% (Suetsugu et al., 2004), 28-56% (Anovitz et al., 2013).

As a result, more broadly, capillarity may play a complex, readily hidden, role in how pore structures change when exposed to wetting–drying cycles or are subject to increasing aridity conditions. Global climate change with more frequent extreme events (e.g. Destouni and Verrot, 2014; Touhami et al., 2015; Wanders and Wada, 2015) increases the extent of arid cold and warm areas, possible hosts of the behaviors presented here. The anthropogenic interests for deep geological storage

targeted complex sedimentary formations capped by thin porous rocks, which are typically prone to strong capillarity states.

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