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# Nanoplastics identification in complex environmental matrices: Strategies for polystyrene and polypropylene

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## Abstract

Identification of nanoplastics in complex environmental matrices remains a challenge. Despite the increase in nanoplastics studies, there is a lack of studies dedicated to nanoplastics detection, partially explained by their carbon-based structure, their wide variety of composition, and their low environmental concentrations compared to the natural organic matter. Here, pyrolysis coupled to a GCMS instrumental setup provided a relevant analytical response for polypropylene and polystyrene nanoplastic suspensions. Specific pyrolysis markers and their indicative fragment ions were selected and validated. Possible interferences with environmental matrices were explored by spiking nanoplastics in various organic matter suspensions (i.e., algae, soil natural organic matter, and soil humic acid) and analyzing an environmental suspension of nanoplastics. While a rapid polypropylene nanoplastics identification was validated, polystyrene nanoplastics require preliminary treatment. The here presented strategies open new possibilities to the detection/identification of nanoplastics in environmental matrices such as soil, dust, and biota.

## Keywords

Nanoplastics, polypropylene, Py-GCMS, identification, OM interference, complex matrices.

## Introduction

Under environmental conditions, all plastic debris is likely to release many nanoplastics (NPs) under multiple physicochemical degradation pathways<sup>1-3</sup>. While the scientific community is now admitting the occurrence of NPs in the environment, relatively few studies demonstrated their presence<sup>4-6</sup>. When the material reaches the nanoscale, its properties, reactivity, and impact are significantly enhanced, especially for plastic debris due to its considerable heterogeneity in composition<sup>7-9</sup>. The plastic debris regroups an extensive range of materials which complexifies the analytical and environmental challenge. This challenge is even more complex when the size of the plastic debris is under the optical resolution<sup>8,9</sup>. Each plastic family, composition, and size require specific attention and dedicated strategy as performed this last century for a wide range of contaminants.

Among the various plastics produced globally, polypropylene and polystyrene represent approximately 20 to 30% of the total production<sup>10</sup>. Polypropylene plastic debris are detected in diverse environmental compartments<sup>10,11</sup>. For instance, Reisser et al.<sup>12</sup> found approximately 30% polypropylene in plastic debris collected on the Australian sea coast. Brignac et al.<sup>13</sup> consistently found a significant proportion of polypropylene pieces (i.e., >15%) on beaches, sea surface, and seafloor. In terrestrial environments, polypropylene microplastics (MPs) have been reported in sewage sludge<sup>14</sup>, sediments, and suspended matter in rivers<sup>15</sup>. While Polystyrene-based plastics represent more than 96% of the plastic debris investigated in the literature, it represents less than 6% of the plastic found in the environment<sup>16,17</sup>.

At the nanoscale, determining plastic debris occurrence in complex environmental media is highly challenging due to their small size and ability to heteroaggregates with the natural organic matter (NOM) and the inability to physically separate them (NPs vs. NOM) by filtration or other processes<sup>18</sup>. As recently explained, when reaching the nanoscale, the volume of NOM related to the plastics increases exponentially. Therefore, analytical strategies used to characterize microscale cannot be extrapolated to the nanoscale plastic debris.

Among the available analytical methods, pyrolysis coupled to gas chromatography and mass spectrometry proved the presence of polyethylene NPs in the colloidal fraction of the North Atlantic Ocean gyre (NAOG)<sup>4</sup>. However, we could not detect polypropylene markers in the colloid fraction, while this plastic is present in the microplastic fraction. Two reasons were invoked: 1) low polypropylene abundance and 2) modification of the pyrolysis products of nanoscale polypropylene. To date, no PyGCMS pyrogram is provided for colloidal polypropylene in the literature., Potential variations in pyrolysis products at the nanoscale are thus unknown. Polystyrene NPs are identified in NAOG<sup>4</sup>, coast exposed to the NAOG<sup>5</sup>, and contaminated soil<sup>6</sup>. However, as for polypropylene under complex environmental conditions, interferences with OM were encountered, especially for their primary direct markers, styrene, and toluene<sup>19,20</sup>. While authors are looking for the styrene dimers as proof of the polystyrene presence, as observed for a large fragment of polyethylene, it is less probable to found such pyrolyze products at the nanoscale due to the considerable heterogeneity of the NOM. One way to by-pass such a limitation is to investigate the pyrolyzates ratio to detect the presence of the plastics in NOM<sup>21</sup>. Moreover, the problem is even more complicated as, in addition to pyrolyzates interferences, we demonstrated that the NOM also attenuates the global signal<sup>6</sup>. A critical preliminary investigation of the possible interferences with NOM has therefore to be performed.

Therefore, this study aims to explore polystyrene and polypropylene NPs identification in the presence of different NOM by PyGCMS. Two different and complementary strategies are developed and proposed according to the plastics composition and the NOM. Based on a careful

investigation of chromatograms in terms of pyrolyzate and retention time. We demonstrate that polypropylene NPs can be directly identified in complex media. However, for polystyrene NPs, we needed to develop a protocol for purifying the environmental sample from NOM to obtain a reliable detection of polystyrene in complex media. The method is finally applied on an environmentally-relevant NPs suspension obtained from plastic debris altered under environmental conditions and covered by NOM. While we use the pyrolysis library to identify polypropylene NPs, polystyrene NPs are detected in the environmental sample according to the relative ratio of its principal pyrolyzates.

## **Materials and methods**

### Samples

Polystyrene and polypropylene pellets were purchased from Sigma Aldrich and were used as reference materials for PyGCMS and to produce nanoplastics. Natural organic matter standards are purchased from IHSS (International Humic Substances Society, and, respectively): leonardite NOM (LNOM, references 1S104H), leonardite Humic Acid (LHA, reference 1BS104L). LNOM and LHA were solubilized in 1 mol L<sup>-1</sup> of NaOH (Sigma Aldrich) (1:10 ratio in g/g) for 24 h. Additionally, to these NOM standards, fresh organic matter is prepared from Sargasso algae residue (Sainte-Marie Bay, Guadeloupe, France). The organic matter of Sargasso algae (SAWE) was obtained by stirring 1 g of algae in ultrapure water (1:10 ratio in g/g) for 24 h and filtered at 3 µm (Pall)<sup>22</sup>. The total organic carbon concentration (TOC) of all the organic matter solutions was quantified by a TOC-analyzer (TOC-V analyzer, Shimadzu).

### Nanoplastics preparation

Polypropylene and polystyrene NPs are produced using a top-down mechanical degradation as previously described<sup>23</sup>. Briefly, plastic pellets were fragmented using a planetary ball mill (Pulverisette 7, Fritsch GmbH) with a 5 or 10 mm diameter zirconium oxide ball. The fragmented polypropylene and polystyrene powders are freeze-dried (Alpha 3-4, Christ). Then, the powders are redispersed in water and filtered at 3 µm (polyethersulfone membrane, Pall) to remove microsized-particles. Additionally, environmentally relevant nanoplastics, namely environmental NPs (EnvNPs), were produced from plastic debris collected in the environment<sup>24</sup>. Plastic debris were randomly collected on Saint Marie Bay's beach (Guadeloupe, France) in September 2018 (Figure 1)<sup>13</sup>. This beach receives plastic debris and Sargassum Algae (brown line on Fig. 1b and 1c) from the NAOG<sup>5</sup>. Briefly, 100 g of plastic debris was stirred in ultrapure water (1:5 ratio) for 2 days (KS 15A, Edmund Bühler GmbH), sonicated for 5 days (TP 680/DH, Elma GmbH), and finally filtered at 3 µm (PTFE, Pall). To be in the range of analytical ability of the here-used methods and when necessary, the mass of the colloids (analyzed nanoplastics and NOM) was concentrated using an ultrafiltration system (Amicon, Merck Millipore) equipped with a 20 kDa cut-off membrane (PES, Microdyn NADIR). To purify the nanoplastics from the NOM, solutions were oxidized by H<sub>2</sub>O<sub>2</sub> and UV. The UV (UVC) was used within the range of 24h (RMR-600, Rayonet). As beached plastics were primarily associated with sand and sargassum algae residues (see supplemental information, Figure S1), a significant amount of natural organic matter is released and dispersed with the nanoplastics produced. All the final dispersions were kept in the dark and at 4°C and remain stable over 3 weeks.

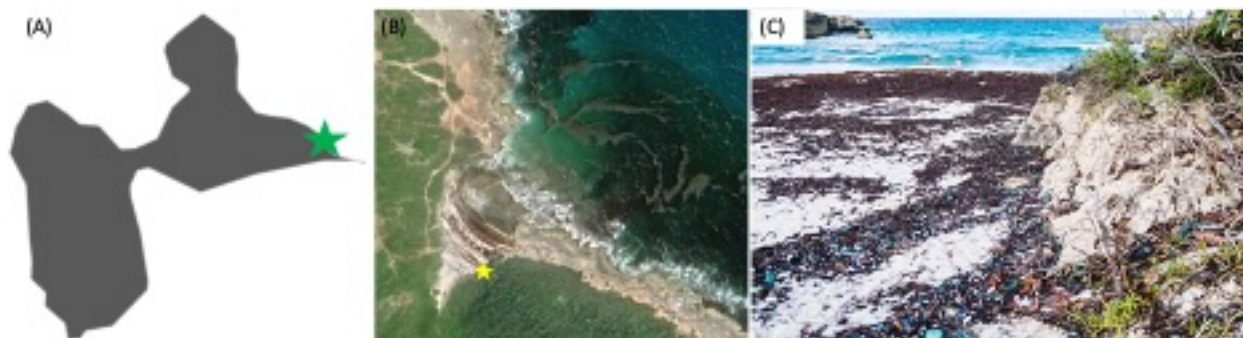


Figure 1: (A) Sampling area on the French Caribbean Island Guadeloupe; (B) localization of the bay (Sainte Marie) at the GPS coordinate 16°16'24" N-61°14'35" W with the brown line is corresponding to the Sargassum Algae coming from the oceans (CNES); (C) pictures of the sampling site where the plastic debris are identified and mixed with the Sargassum Algae (CNRS, Cyril Fresillon).

### Size and shape characterization

The hydrodynamic size distribution of the NPs and EnvNPs was determined using optical fiber dynamic light scattering (DLS). Measurements are performed with a VASCO Flex (Cordouan Technologies, France). The size distribution of NPs is determined by fitting the autocorrelation function (ACF) with Padé-Laplace and Sparse Bayesian Learning algorithm<sup>25,26</sup>. Each DLS analysis is carried out with six replicates with an analysis period of 120 s. Only results for replicates with reliable residuals ( $r < 0.01$ ) were accepted. The ACF samples are compared to 200, 500, and 900 nm polystyrene spherical latex nanoparticles (Thermo Scientific). The shape of the NPs is examined using a Jeol JEM 2100 HR electron microscope in transmission mode (TEM) after drying 2.5  $\mu\text{L}$  of suspension onto a carbon grid (Oxford Instrument). Images are acquired with a Gatan Orius SC200D camera, and elemental analysis is performed using an EDX Oxford X-Max 80T detector.

### Nanoplastics molecular identification

NPs suspensions (Polypropylene, Polystyrene, and EnvNPs) and the NOM were analyzed by pyrolysis (PY-3030, Frontier Lab) coupled to gas chromatography and a mass spectrometry detector (Py-GCMS) (5977B, Agilent Technologies). Before Py-GCMS analysis, liquid samples are evaporated at 45°C in an 80  $\mu\text{L}$  pyrolysis cup (Frontier Lab). The mass of prepared nanoplastics was fixed at 2  $\mu\text{gC}$  in the pyrolysis cup. Concerning the NOM, the mass was ranged from 25 to 175  $\mu\text{gC}$ . To assess any polymer contamination during the sample preparation (i.e., drying process), empty pyrolysis cups were systematically added beside regular samples. The optimal GC conditions were fixed according to the method of Dehaut et al.<sup>27</sup>. The pyrolysis was performed at 600°C. Once pyrolyzed, samples were injected (splitless) into a 30m DB5 capillary column (Agilent Technologies) using helium as the carrier gas. The column temperature was fixed at 50°C for 10 min, increased to 275°C at a rate of 5°C/min, and finally kept at this temperature for 15 min. Separated pyrolysis products were ionized at 70 eV, and their mass spectra were analyzed between  $m/z$  33 to 500. Pyrolyzates signals are compared to the NIST library and personal library builds using plastic pellets described above. Each pyrograms were realized in four replicates. As it is not yet possible to quantify nanoplastics by PyGCMS, we determined the signal-to-noise ratio (SNR),

ranging from 150 to 430, for all performed analyses. Based on the SNR and the mass of nanoplastics<sup>28</sup>, LOD can be estimated from 0.2 to 0.4 mgC L<sup>-1</sup>.

## Results and Discussions

### Nanoscale effects on plastic pyrolysis.

The size and shape of polypropylene and polystyrene NPs used in these studies were described in the supplemental information (Figure S2). For both suspensions, NPs present a different population with  $d_H < 1 \mu\text{m}$ . It is worth noting that hydrodynamic diameter depends on the geometric size (radius) and the shape of materials<sup>29,30</sup>. Therefore, asymmetrical materials can present different diffusion (transversal and rotational), which results in different apparent populations on the size distribution<sup>30</sup>. All NPs suspensions present anisotropic and heterogeneous shapes with spherical equivalent sizes less than 1  $\mu\text{m}$ . For polypropylene NPs, fine particles (size > 200nm) were identified as NPs, and opaque particles as inorganic species (< 200nm), as confirmed by EDX analysis (Figure S3). These species are principally titanium oxide, known to be one of the major plastic additives<sup>31</sup>. No inorganic species were identified into polystyrene NPs, which is explained by their virgin pellet source. Based on the polypropylene pyrolysis at a microscale (see SI, Figure S4 and Table S1), four distinctive peaks are pre-selected: the C9 (2,4-dimethyl-1-heptene), the C12 (meso-2,4,6-trimethyl-1-nonene), the C15i (isotactic-2,4,6,8-tetramethyl-1-undecene) and the C15s (syndiotactic-2,4,6,8-tetramethyl-1-undecene). The pyrolyzates C9 and the C15s have already been documented as a polypropylene marker in environmental studies<sup>4,15,14,32,33</sup>. Based on a literature comparison<sup>14,34</sup>, none of these markers were obtained for various NOMs. The pyrograms of polypropylene NPs (TIC,  $m/z$  70 and 69) are plotted in Figure 2. The  $m/z$  70 pyrogram specifically enhanced the marker C9, whereas the  $m/z$  69 enhanced the other polypropylene pyrolyzates (i.e., C10, C12, C15i, and C15s). Peaks are observed at the retention times of all the selected polypropylene markers (Figure 2A). The mass spectrum of polypropylene NPs obtained at the characteristic retention times (blue bar) were compared with the polypropylene library (red bar). For C9, C12, and both C15 compounds, mass spectra were significantly similar to the library ones (Spectrum Similarity,  $SS > 0.90$ ). Although these markers are found on both micrometric pellet and NPs pyrolysis, their relative peak areas vary significantly (i.e., Table S1). For polypropylene NPs, the proportion of C12, C15i, and C15s increases by a factor of 2-3 compared to MPs pellets. Based on this result, C9, C12, and C15 compounds are the most relevant for identifying polypropylene at the nanoscale.

For polystyrene NPs identification, seven distinctive peaks were pre-selected from the microscale pellet pyrolysis (see SI, Fig. S5): toluene, styrene,  $\alpha$ -methyl styrene, bibenzyl, styrene dimer, 2-phenyl naphthalene, and styrene trimer. Only two pyrolyzates differ from the pyrolysis of literature and hereafter studied organic matter (Table S2): styrene dimer and trimer. The  $m/z$  104 pyrogram specifically enhanced the styrene, whereas the  $m/z$  91 pyrogram enhanced the styrene dimer and trimer. Only significant peaks at the styrene and styrene dimer's retention time were studied (see SI, Fig. S6). Their mass spectra were similar to the library spectra. Oppositely, the styrene trimer mass spectra are noisily due to their low abundance and the limit of detection. From polystyrene MPs to NPs, the proportion of styrene dimer and trimer decreased from 26% to 0.2% and 23% to a negligible area. One hypothesis to explain such result, as we previously observed for polyethylene<sup>4</sup>, is that the probability of forming a large mass of pyrolyzate (i.e.,  $m/z > 200$ ) considerably decreases from the microscale to the nanoscale.

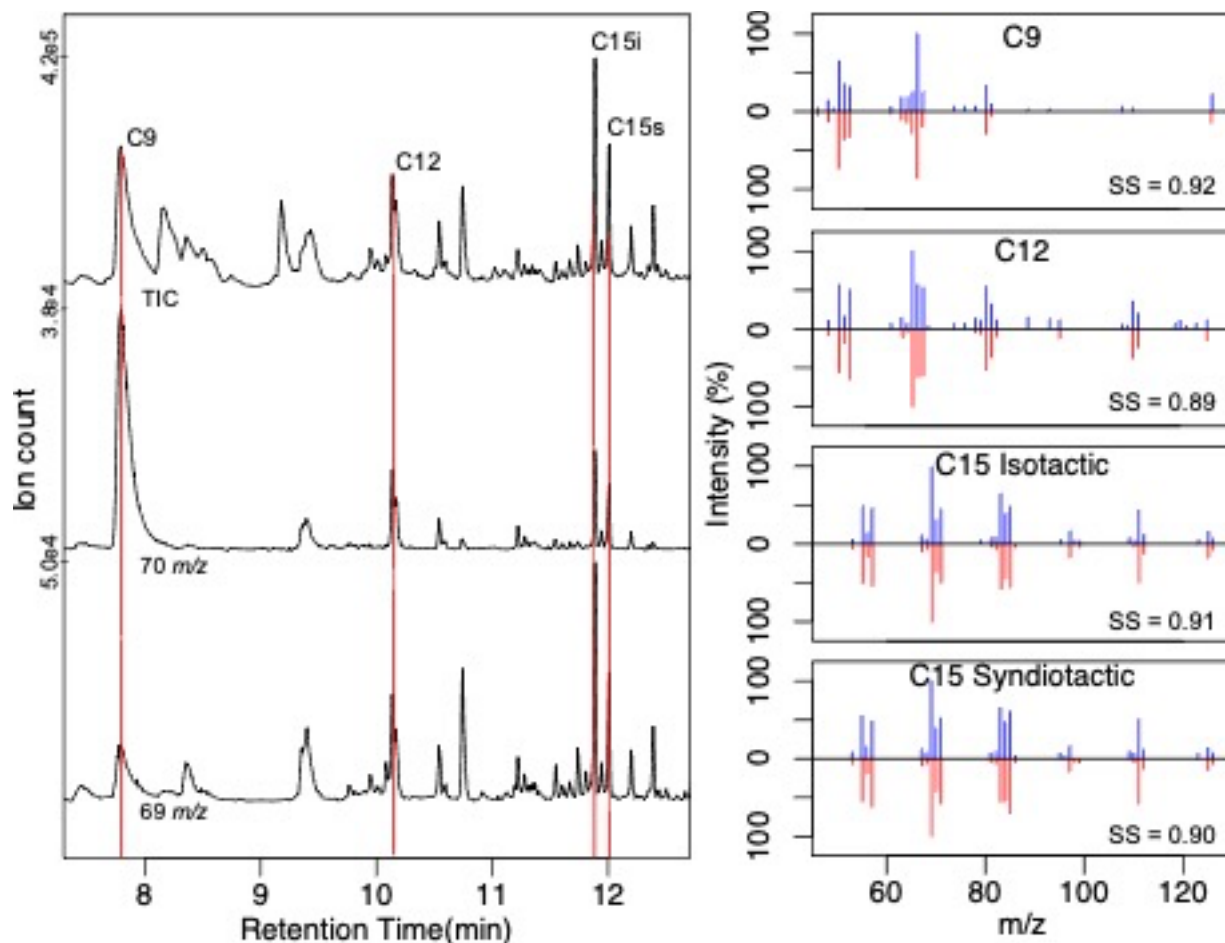


Figure 2: (A) polypropylene NPs pyrograms for all TIC,  $m/z$  70 and  $m/z$  69. (B) Mass spectra comparison between the MSs of pyrolyzates recorded at the PP marker retention times (blue) and their library MSs (red).

Since the direct identification of polystyrene NPs is based on both pyrolyzates, the decreasing size from polystyrene MPs to NPs would affect the polystyrene NPs detection in environmental matrices. Regarding their nanoscale, low quantity, and ability to hetero-aggregate with the NOM, the direct identification of polystyrene NPs in environmental matrices seems impossible. Polystyrene identification should thus integrate an indirect identification. An alternative is to determine the absolute pyrolyzate ratio of toluene/styrene<sup>35,21</sup>. This ratio is  $<0.01$  for pristine polystyrene and  $> 4-5$  for organic matter<sup>21,35</sup>.

#### Natural organic matter contribution

Although few pyrolyzates are common between polypropylene NPs and NOM, their potential coelution of NOM and polypropylene markers was thus investigating. Polypropylene identification was performed by spiking polypropylene NPs into the three different OM (Materials and Methods section). For the SAWE, the polypropylene/OM ratio (g/g) was set to 0.073, corresponding to  $2 \mu\text{gC}$  of polypropylene and  $27.5 \mu\text{gC}$  of SAWE. For LNOM and LHA, the polypropylene/OM ratio (g/g) was set to 0.011, corresponding to  $2 \mu\text{gC}$  of polypropylene and  $175 \mu\text{gC}$  of OM. Leonardite organic matter pyrograms with and without polypropylene NPs were compared to identify potential coeluted pyrolyzates at the retention times of C9, C12, and both C15 compounds.

For polypropylene NPs spiked into SAWE, at  $m/z$  70 (Figure 3A), the pyrogram showed a significant peak at the retention time of C9. At  $m/z$  69, the three other polypropylene markers are observed in the pyrogram (Figure 3B). Without polypropylene NPs, the mass spectra did not show any similarity with the reference, confirming that SAWE and polypropylene have no homo-pyrolyzate interferences, as previously discussed (Table S1). The mass spectra for polypropylene NPs spiked into SAWE showed a very high similarity with polypropylene for the following markers: C9, C15i, and C15s (Figure 3C). These coeluted SAWE pyrolyzates were thus at too low concentration to interfere. Conversely, the C12 spectra were not similar due to the presence of a coeluted OM pyrolyzates, which contributed to the C12 mass spectrum (Figure 3C). Therefore, only C12 has coeluted-pyrolyzate interference with SAWE.

For polypropylene NPs spiked into LNOM (Fig S7), the  $m/z$  70 and  $m/z$  69 pyrograms showed a significant peak increase for all the polypropylene markers compared to non-spiked LNOM. The peaks at polypropylene retention times remained low compared to those of OM. The LNOM mass spectra with and without polypropylene NPs were different. No homo-pyrolyzate interferences are observed between LNOM and polypropylene. For LHA, a significant peak increase was observed for pyrograms with and without polypropylene NPs at all polypropylene markers' retention times (Fig S8). The polypropylene markers are thus very distinctive. No mass spectrum similarities are obtained between polypropylene markers and LHA. As observed for SAWE and NOM, no homo-pyrolyzate interference exists between LHA and polypropylene. With polypropylene NPs, the mass spectra were similar for the polypropylene markers (Figure 3C). No coeluted-pyrolyzate is observed. Thus, polypropylene can be directly identified with all four selected markers.

In summary, coeluted-pyrolyzate interference is identified for C12 for SAWE and NOM. Polypropylene was still identifiable because the three other markers had no interferences. The use of these three polypropylene markers allowed direct polypropylene identification in the three studied OM samples. Concerning the environmental identification of polypropylene, although OM was in excess (e.g., 14 times for SAWE and 100 for LNOM and LHA), all these ratios allowed polypropylene identification. Moreover, the  $m/z$  70 pyrograms show that these ratios could be decreased and that polypropylene identification could be achieved using only the C9 marker. For LNOM and LHA, the polypropylene/OM ratio could be decreased since no peak is present at the retention time of C9 in both OM pyrogram without polypropylene NPs.

In contrast, the SAWE pyrogram without polypropylene NPs exhibited a peak at the retention time of C9. Therefore, over the polypropylene-OM ratio, the C9 marker could no longer be identified. By setting the limit when the C9 mass peak is equal to 2 times the SAWE coeluted pyrolyzate, it may be difficult to assess the presence of polypropylene in SAWE below a ratio of 0.025 (g/g).



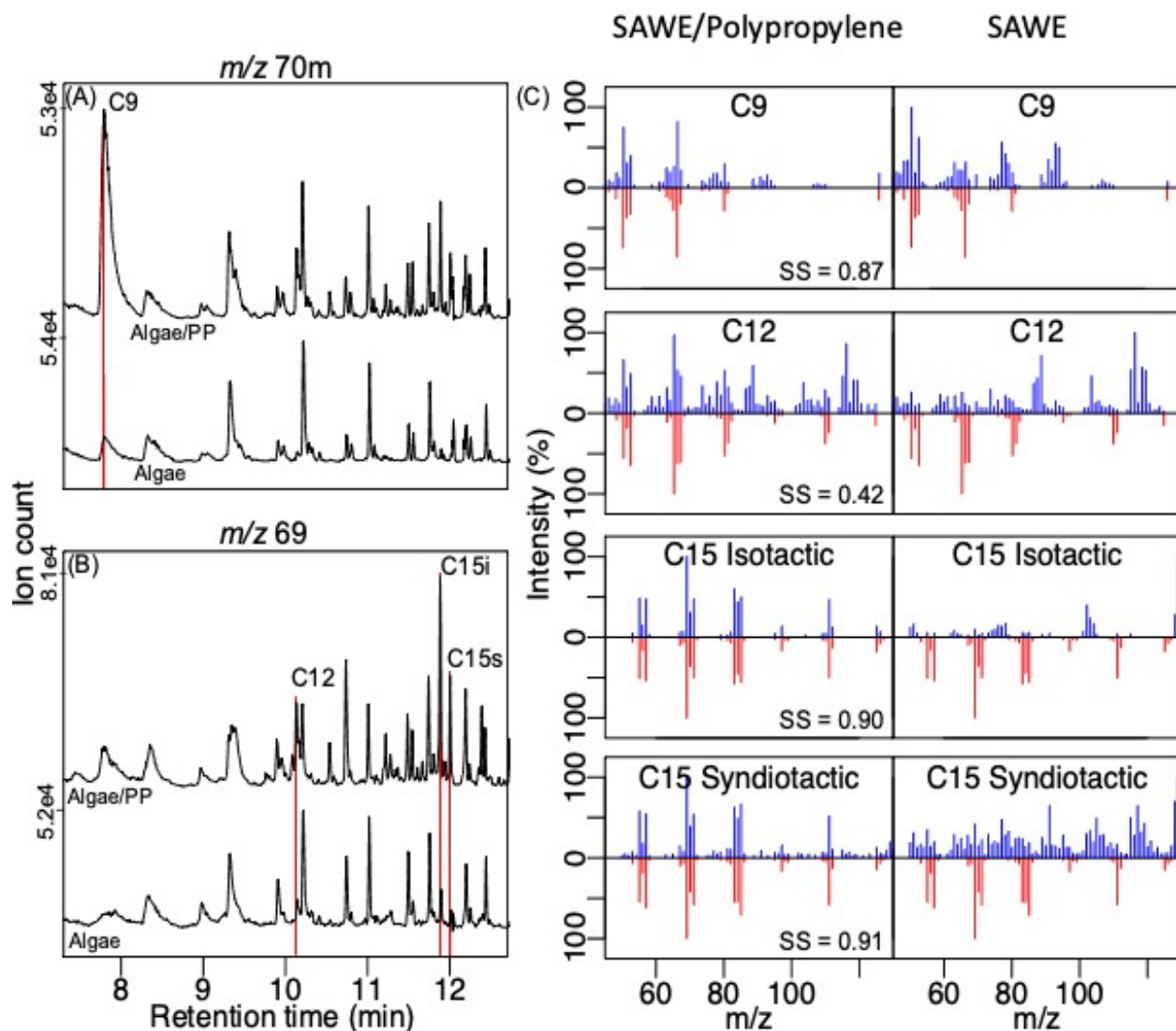


Figure 3: (A)  $m/z$  70 pyrograms for SAWE without and with polypropylene NPs; (B)  $m/z$  69 pyrograms for SAWE without and with polypropylene NPs; (C) Mass spectra of pyrolyzates recorded at polypropylene marker retention times (blue) and their library MSs (red).

#### Application to environmental nanoplastics

To evaluate the ability to identify the NPs composition in complex matrices, we applied the PyGCMS on environmental nanoplastics (EnvNPs). These EnvNPs were produced from unwashed beached plastics covered by Sargasso algae (e.g., type SAWE). Figure 4 illustrates the size and shape characterization of the EnvNPs obtained by DLS and TEM. The EnvNPs have a hydrodynamic diameter ranging from 200 to 500 nm (Figure 4), lower than the polystyrene and polypropylene nanoplastics. This difference can be described by the advanced oxidation state of the plastic debris sampled in the environment that more likely produces nanoscale particles. The EnvNPs are less size-polydispersed and better separated probably because NOM stabilized them differently according to their size. However, EnvNPs present an identical shape to polypropylene and polystyrene NPs (Figure 4). As for polypropylene NPs,  $\text{TiO}_2$  and carbonates were identified.

While TiO<sub>2</sub> results from additive, carbonates could result from both additive and natural biotic or abiotic precipitates.

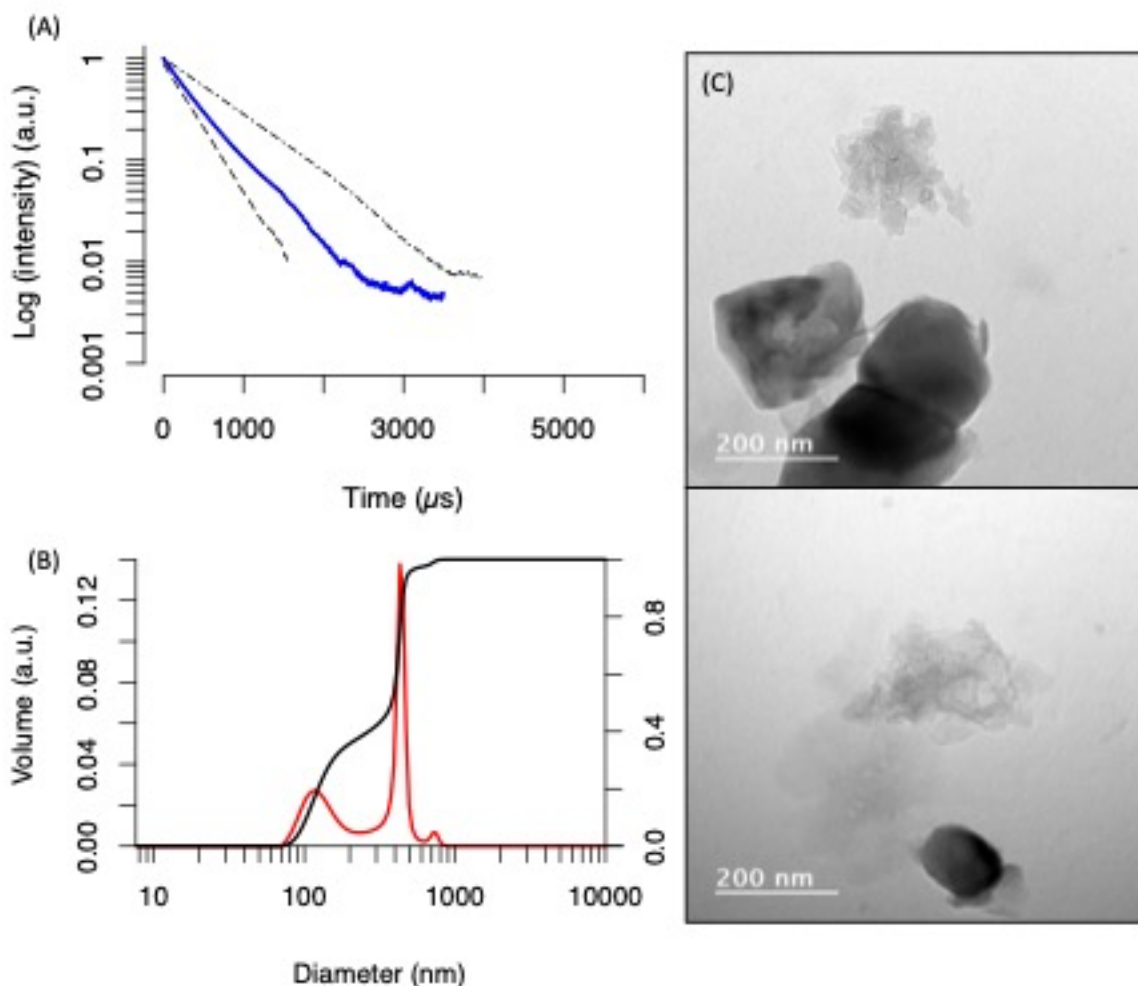


Figure 4: (A) ACF function of the EnvNPs bordered by the 500 and 900 nm polystyrene spherical latex nanoparticles; (B) corresponding size distribution expressed in volume with the associated integration; (C) TEM images of EnvNPs.

The direct identification of polypropylene NPs was applied and validated with EnvNPs. For EnvNPs, on the  $m/z$  70 pyrogram, an important peak occurred at the retention time of C9 (SI, Fig. S9). On the  $m/z$  69 pyrogram, three other high-intensity peaks are distinct at the polypropylene markers' retention time: C12, C15i, and C15s. For all of these peaks, comparisons of their mass spectra with those in the literature showed high similarities ( $SS > 0.90$ , Figure S9). Therefore, even if NOM is highly present in the EnvNPs suspension, polypropylene is evidenced without further treatments.

However, concerning Polystyrene in the EnvNPs, identifying the styrene dimer and trimer pyrolyzates was not possible, as illustrated by the pyrograms in Figure 5 (Fig. 5A). Moreover, based on the TIC, a large range of signals and peaks are NOM characteristics without a clear indication of polystyrene. Also, the Tol/Sty ratio obtained for EnvNPs ranges from 4 to 5, suggesting that the Styrene signal ( $m/z$  104) can be attributed to NOM rather than polystyrene.

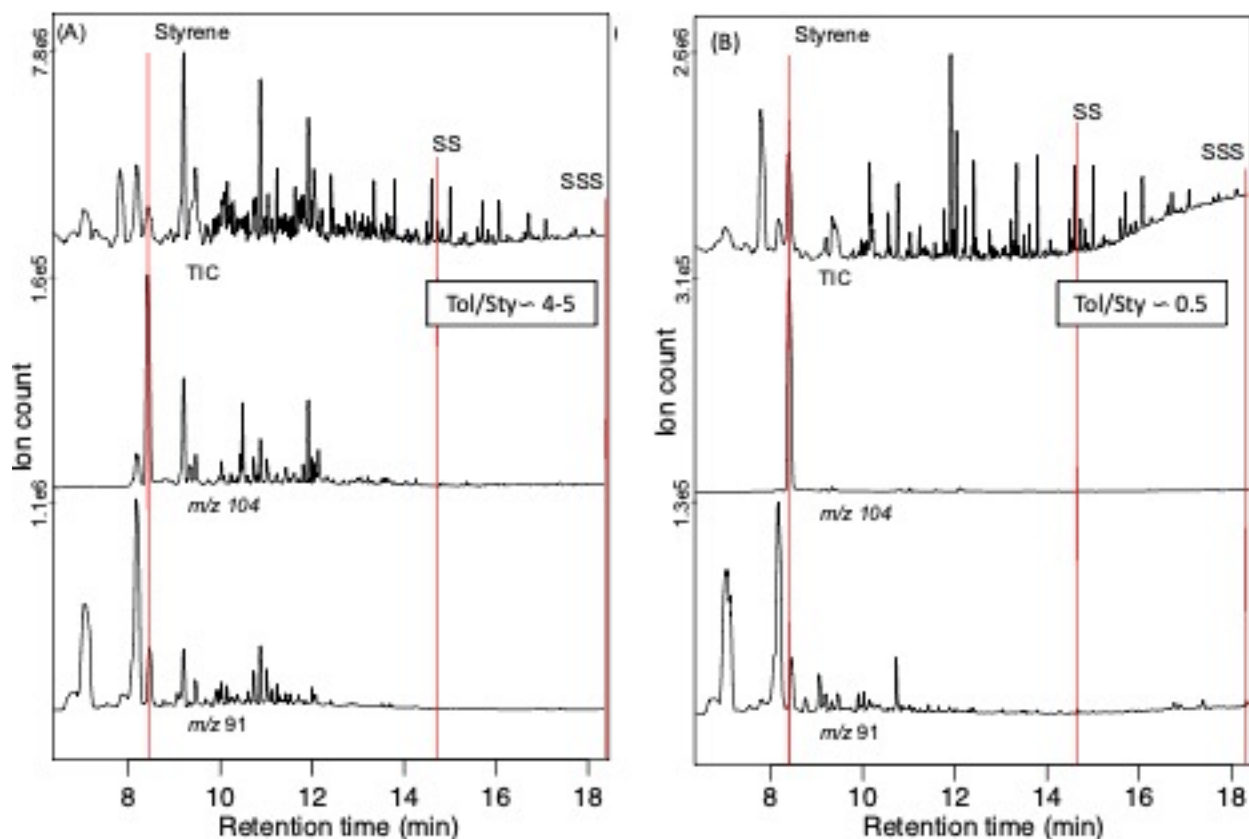


Figure 5: EnvNPs pyrograms for TIC,  $m/z$  104 and  $m/z$  91 (A) before and (B) after  $H_2O_2/UV$  purification process.

Therefore, we developed a protocol using  $H_2O_2$  and UV light<sup>36,37</sup> to selectively degrade NOM without affecting the nanoplastics integrity by testing various experimental conditions. The optimized protocol consists of the  $H_2O_2$  addition at 1% (v/v) in the EnvNPs suspension and then exposition to UV light to induce OH radical reaction. This protocol allows the degrading of more than 90% of NOM (SI, Fig S10). During the  $H_2O_2/UV$  processes, the Tol/Sty ratio is followed over time and NOM degradation (Figure 6). As explained above, at  $t=0$ , EnvNPs had a high Tol/Sty ratio suggesting that the styrene comes from NOM as expected by the large amount of NOM associated with the plastics debris<sup>35,38</sup>. After twenty-four hours of UV, the Tol/Sty ratio decreases to reach a ratio around 0.5, highly similar to the polystyrene NPs' ratio (Fig. 5B). It is worth noting that after EnvNPs purification from NOM, the  $m/z$  104 pyrogram presents both a cleaner and twice intense chromatogram with a clear styrene peak ( $t_R=8.2$  min). At the same time, the carbon mass significantly decreased with the  $H_2O_2/UV$  light action. Such combined observation confirms, after all, the polystyrene occurrence in the EnvNPs suspension.

Additionally, the detection of the styrene is significantly enhanced after the  $H_2O_2/UV$  treatment. This result is coherent with: (i) the better sensitivity of the PyGCMS to detect styrene pyrolyzate for polystyrene vs. NOM since styrene is the principal component of polystyrene; (ii) the NOM effect on the signal attenuation previously demonstrated for styrene for Nps detection in soil<sup>6</sup>. Moreover, the major pyrolyzate of polypropylene (i.e., C9, C12, C15i, and s) were still identified (SI, Fig. S11). The combination of both strategies is thus appropriate to detect them in complex environmental matrices.

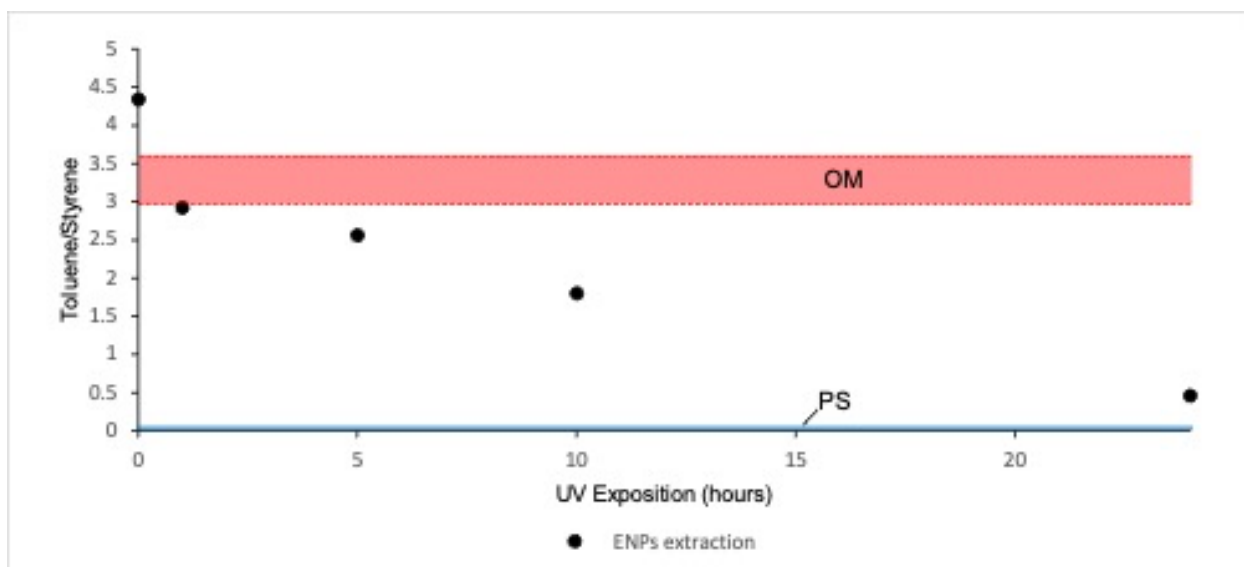


Figure 6: Toluene/Styrene ratio evolution during the OM degradation applied on the EnvNPs suspension. The red zone indicated the toluene/styrene ratio obtained to the different studied NOM. The blue zone indicated the toluene/styrene ratio obtained for the polystyrene microscale pellet and polystyrene NPs.

## Supporting Information

The following files are available free of charge. An example of plastic-coated on Sargasso algae (Figure S1); Size description of the nanoplastic models (Fig. S2). TEM image and EDX spectra from a nanoplastic and the TEM grid (Figure S3). Polypropylene and polystyrene pre-selected pyrolyzates and peak information obtained from both micro and nanoscale plastics pyrolysis (Fig. S4, S5, S6, and Table S1 S2). Standards and Environmental nanoplastics pyrograms with and without natural organic matters (Fig. S7, S8, S9, S11). Evaluation of the degradation efficiency of the natural organic matter (Fig. S10).

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## Notes

There are no conflicts to declare.

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## Abbreviations

C9, 2,4-dimethyl-1-heptene; C10, 2,4,6-trimethyl-1-heptene; C12, meso-2,4,6-trimethyl-1-nonene; C15i, isotactic-2,4,6,8-tetramethyl-1-undecene; C15s, syndiotactic-2,4,6,8-tetramethyl-1-undecene.

## Synopsis

Single out the NANOPLASTICS in environmental matrices is still challenging; this study provides evidence of Nanoplastics direct identification by Py-GCMS.

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